Photophysical and photochemical studies of tricarbonyl rhenium(I) N-heterocyclic carbene complexes containing azide and triazolate ligands†

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Rhenium(I) N-heterocyclic carbene (NHC) complexes of the type fac-\([\text{Re}(\text{CO})_3(NHC)]\) with either azide or triazole ancillary ligands L and pyridyl or pyrimidyl substituted imidazolyl units have been prepared and structurally characterised, and their photophysical and photochemical properties studied. All of the complexes exhibit phosphorescent emission from triplet metal-to-ligand (3MCLT) excited states, typical of triacarbonyl Re(I) complexes, with the triazolate bound complexes having higher quantum yields and longer decay lifetimes compared to the azide bound complexes. The complexes containing pyridyl substituted imidazolyl units are photoactive when dissolved in acetonitrile and undergo photochemical CO dissociation, the rate of which is significantly greater in the azide cf. triazolate complex. The photochemical mechanism of the azide/pyridyl complex was analysed and appears to give the same products, albeit with different ratios, to previously reported complexes where L is a halide. A reaction mechanism is proposed.

Introduction

Rhenium(I) complexes of the type fac-\([\text{Re}(\text{CO})_3(diim)L]\)^{0+}, where diim represents a π-conjugated diimine ligand such as 2,2′-bipyridine (bpy) or 1,10-phenanthroline (phen) and L is a monodentate anionic or neutral ancillary ligand, have been extensively studied due to their rich photophysical properties that give rise to phosphorescent decay typically from their lowest lying triplet metal-to-ligand charge transfer excited state (3MLCT).^{1–3} The emission from these excited states, which often displays relatively high quantum yields and long excited state lifetimes, can be fine tuned by judicious variation of the diim and ancillary ligands, making this class of compounds suitable for many applications including light emitting devices^{4,5} and biological labels.^{6–8}

Although complexes based on diim ligands have been thoroughly investigated, analogous complexes of the type fac-\([\text{Re}(\text{CO})_3(NHC)L)]^{0+}\), where NHC represents a bidentate N-heterocyclic carbene ligand, which binds the metal through an imine type N atom and the carbene C atom, have been considerably less studied despite their ubiquity in inorganic chemistry. Nevertheless, complexes possessing NHC ligands bound to Ru(Ⅱ), Pt(Ⅱ), Ir(Ⅲ) and Au(Ⅰ)8–12 (just to name a few) display rich luminescent properties and so their effect on the photophysical properties of Re(Ⅰ) complexes is of great interest. Along these lines, we13–15 and others16,17 have investigated these types of complexes, where the NHC is substituted in the 1-position with different imine type N donors that bind to the rhenium and thus change their photophysical properties. We have found that if the NHC contains a pyridyl as the N donor, the complex can exhibit photochemical CO dissociation in acetonitrile solution to form a tricarbonyl acetonitrile solvated complex and two different dicarbonyl complexes. While we have modified the N donor group to pyrimidyl, quinolyl, and quinoxyl to probe the scope of the photochemistry, we have found that only the pyridyl containing complexes are photoactive. These photoreactive rhenium complexes could be of potential benefit in the area of photoinducible CO-releasing molecules (photoCORMs), whereby bioactive CO is released from the metal centre in a spatially and temporally controlled manner.18–21

We have recently become interested in the coordination of azide ligands to metal NHC complexes and their associated reactivity toward alkynes in a 1,3-diploar cycloaddition reaction. This area of chemistry has been extensively investigated by the group of Veige,22–25 who have studied the reaction of metal azides and metal alkynyl complexes, and others who have investigated the reactions of metal azides with organic alkynes.26–28 In the

**NJC**

PAPER

Cite this: New J. Chem., 2016, 40, 5797

Received (in Victoria, Australia) 22nd November 2015, Accepted 21st February 2016

DOI: 10.1039/c5nj03301b

Electronic supplementary information (ESI) available. CCDC 1431845–1431848. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5nj03301b

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context of CORMs, the inclusion of a reactive azido ligand into the rhenium complex could be used to prepare and study photoactive CO-releasers bound to biological molecules, through an in situ click reaction with alkynyl functionalised biomolecules, without the need to add any additional catalysts such as copper(n) sulfate and ascorbate. Thus, in this study, we aim to expand our knowledge of complexes of the type \( \text{fac-}[\text{Re(CO)}_3(\text{NHC})_2] \) by modifying the L ancillary ligand from exclusively halide (Cl\(^-\) and Br\(^-\)) to azide, and then reacted the metal azide with the model alkyne dimethyl acetylenedicarboxylate to form a triazole ligand via the 1,3-dipolar cycloaddition. Two sets of complexes containing pyridyl and pyrimidyl substituents on the NHC ligands as well as azide and triazole ancillary ligands have been prepared, and their photophysical and photochemical properties studied.

**Results and discussion**

**Synthesis and characterisation**

Treatment of 1a and 1b with silver triflate in acetone and subsequent stirring over an excess of sodium azide led to the formation of 2a and 2b (Scheme 1), which were isolated in yields of 59 and 68%, respectively. The two complexes are soluble in most common organic solvents including dichloromethane, chloroform, acetone, methanol, dimethylsulfoxide, and dimethylformamide, and even exhibit moderate solubility in methane, chloroform, acetone, methanol, dimethylsulfoxide, and dimethylformamide. The disappearance of the azide stretch in the IR spectra of 3a and 3b, which were isolated as off-white solids in yields of 52 and 63%, respectively. Complexes 3a and 3b display drastically reduced solubility in organic solvents and are moderately soluble in acetone, methanol, dimethylsulfoxide, and dimethylformamide. The disappearance of the azide stretch in the IR spectra of 3a and 3b confirms the consumption of the azide precursors, while the presence of three stretches in the 2020–1870 cm\(^{-1}\) region is consistent for the facially arranged CO ligands. The \(^1\)H NMR spectra of 3a and 3b is not straightforward: for every signal in the \(^1\)H NMR spectrum of the triazole methyl singlet, which instead shows two smaller signals of half integration. Analysis of the \(^1\)H NMR spectra reveals two isomers in a ratio of 1 : 0.18, which are ascribed to the coordination of the triazole ligand via the N(2) and N(1) nitrogen atoms, respectively. These isomers could not be separated. As demonstrated by Veige and co-workers, the cycloaddition reaction between gold or platinum azides and gold alkynyl complexes will form N(1) coordinated triazolates.

In the case of octahedral metal complexes, however, steric interactions between the ligands and the newly formed triazole will be such that the kinetically formed N(1) coordinated triazole will favour isomerisation to the N(2) form. Although single crystals of 3a and 3b could be obtained (see X-ray studies), we could not produce a sufficient quantity of crystals that would allow for a detailed NMR study of the interconversion of the triazole ligand to be performed. Interestingly, coordination of a triazole ligand via the N(1) nitrogen has only been reported twice before in the ruthenium acetylocetonate (acac) complexes 4–6\(^{19}\) and also in hexanuclear rhenium chalcogenide cluster 8, see below in Fig. 1.\(^{30}\) In these cases, where two methyl esters are on the backbone of the N(1) bound triazole ligand, the methyl group are seen as two separate singlets in the \(^1\)H NMR spectra in a similar fashion to 3a and 3b. In the case of the ruthenium acetylocetonate (acac\(^-\)) complexes, the formation of the N(1) or N(2) bound triazole was heavily dependent on the

![Image](https://source.unsplash.com/random/300x200)

**Scheme 1** Synthesis of azido and triazole complexes via a 1,3-dipolar cycloaddition.
substituents on the triazole and acac− ligands, with the complexes 5–7 being formed as sole isomer formed in the reaction. Rhenium cluster 8, however, was formed in only ca. 2% relative to the N(2) bound major isomer.

X-Ray structural investigation

Single crystals suitable for X-ray diffraction studies were grown for each of the complexes (see Experimental section for crystal growth conditions), projections of which can be seen in Fig. 2. The structures are consistent with the proposed formulations of the complexes, with each possessing facially arranged tricarbonyl ligands. In all cases, the plane of the mesityl substituent is oriented ca. 90° with respect to the imidazolyl plane. A point of interest within the four structures is the angle between the planes of the imidazole and pyridyl/pyrimidyl rings; while the azido complex 2a is strained in this respect with an interplanar angle of 11.4°, the imidazole unit in complex 2b is almost completely coplanar with the pyrimidyl group, with an interplanar angle of just 1.6°. This seems to be due to the orientation of the azide ligand; in 2a the azide points back toward the carbene, thus bringing it close to one of the methyl groups on the mesityl substituent. The azido ligand in 2b, however, is directed away from the carbene and has no effect on the imidazolyl–pyrimidyl interplanar angle. The rhenium–carbene carbon distances are 2.123(4) and 2.118(5) Å for 2a and 2b, respectively. The nitrogen atoms in the triazole are now slightly further apart from the mesityl methyl group than in the case of 2a and 2b, with distances ranging from 3.76–3.85 and 3.60–3.85 Å for 3a and 3b, respectively. The differences between 3a and 3b can be explained by two factors: the smaller interplanar angle between the triazolate and the Re–C(2)–N(42)–C(101)–C(102) coordination planes in 3b cf. 3a of 52.15° and 57.75°, respectively; and the slightly more twisted nature of the mesityl ligand in 3b, which brings the mesityl methyl group closer to the triazolate. A close examination of the mesityl methyl signals in the 1H NMR spectra of 3a and 3b suggests that these slight differences are maintained in solution, presumably due to the rigidity of these complexes. The aromatic nature of the triazolate ring causes an upfield chemical shift in the 1H NMR spectra of 3a and 3b, with the closer contacts in 3b leading to a slightly more upfield shift of 1.47 ppm, compared to 1.56 ppm for 3a.

Photophysical investigation

The photophysical data of all the complexes in dilute dichloromethane solutions are summarised in Table 1, with all absorption and emissions profiles shown in Fig. 3 (see ESI† Fig. S1 for excitation spectra). The absorption spectra of 2a,b and 3a,b exhibit high-energy bands from ca. 250–300 nm, and lower energy bands of relatively smaller molar absorptivity above 300 nm. The higher energy bands are associated with ligand-centred (LC) π → π* transitions within the NHC and, in the case of 3a and 3b, the triazole ligands. On the other hand, the lower energy bands are assigned to an admixture of MLCT (Re → NHC) and ligand-to-ligand (azide/triazolate → NHC; LLCT) charge transfer transitions, as previously described for similar complexes.13–15 A slight red-shift in the mixed MLCT/LLCT transitions of 2b cf. 2a and 3b cf. 3a of 5 and 6 nm, respectively, is attributed to the more electron deficient pyrimidyl- versus pyridyl-substituted NHC ligands.
Comparison of the maxima of the lower-energy bands in the four complexes shows a decrease of ca. 30 nm going from azide- to triazolate-ligated species, which can be attributed to the electron-withdrawing nature of the triazolate ligand that serves to lower the energy of the HOMO, which is located primarily on the metal centre.

The emission profiles of the four complexes show broad and structureless bands located at 500–650 nm, which are common for emissive states of MLCT/LLCT character. Complexes 2a and 2b, possessing the azide ancillary ligands, have lower energy emissions than 3a and 3b due to the electron-donating ability of the azide ligand raising the energy of the HOMO. There is a general red-shift in the emission from the pyrimidyl-based complexes relative to their pyridyl-based analogues, which is particularly prevalent for the azide complexes 2a and 2b, where the $\lambda_{em}$ are 564 and 645 nm, respectively. This difference is smaller in the triazolate containing complexes, where the $\lambda_{em}$ for 3a and 3b are located at 504 and 559 nm, respectively. The emission wavelengths of 2a,b and 3a,b are significantly different than previously reported analogous complexes containing halide ancillary ligands, which suggests that the presence of either an azide or triazolate ligand strongly influences the energy of the excited states. The lifetime decays ($\tau$) of the complexes are mono-exponential in nature and range from 8 to 115 ns, which then elongate after degassing; a similar trend is observed for the quantum yields, suggestive of an emissive state of triplet multiplicity ($^3$MLCT,$^3$LLCT). The marked differences in the lifetime decays of 2b can be explained, in part, by the energy gap law. Complex 2b, containing both the HOMO-raising azide ligand and the LUMO-lowering pyrimidyl substituent, has a relatively small HOMO–LUMO gap and therefore a high non-radiative decay constant ($k_{nr}$) in the order of $10^8$ s$^{-1}$, leading to relatively fast decay lifetimes and a quantum yield ca. 0.1%. Both triazolate-containing complexes have higher decay lifetimes and quantum yields that their azide-containing analogues. In a frozen matrix at 77 K (see ESI† Fig. S2) all complexes display blue-shifted emission due to the rigidochromic effect, with decay lifetimes ranging from 0.5–8 μs.

### Computational studies

To validate the interpretation of the photophysical data, time-dependent density functional theory (TD-DFT; see Fig. 5 and ESI† for orbital contours and table of calculated transitions) was used to simulate the energetics and absorption spectra of the four complexes. The structures were relaxed in dichloromethane and found to be consistent with the structures obtained from X-ray diffraction. The calculations indicate that the lower energy transitions involve the HOMO–n ($n = 0, 1$) orbitals and the LUMO+m ($m = 0, 1$) orbitals for 2a and 2b; and from the HOMO–n ($n = 0, 1, 2$) orbitals to the LUMO+m ($m = 0, 1$) orbitals for 3a and 3b. In agreement with the band assignment proposed above, the orbital contours show that these transitions are of MLCT nature that are partially mixed with LLCT character, a representative of which is shown in Fig. 4 for complex 2a.

### Table 1 Photophysical properties of the complexes from diluted dichloromethane solutions (ca. 10$^{-5}$ M)

<table>
<thead>
<tr>
<th>Complex</th>
<th>$\lambda_{abs}$ (nm) ($10^4 \varepsilon$ (cm$^{-1}$ M$^{-1}$))</th>
<th>$\lambda_{em}$ (nm)</th>
<th>$\tau^a$ (ns)</th>
<th>$\tau^b$ (ns)</th>
<th>$\phi^{a}$</th>
<th>$\phi^{b}$</th>
<th>$\lambda_{em}$ (nm)</th>
<th>$\tau$ (μs)</th>
</tr>
</thead>
<tbody>
<tr>
<td>2a</td>
<td>228 (3.31), 362 (0.42)</td>
<td>564</td>
<td>45</td>
<td>109</td>
<td>0.007</td>
<td>0.018</td>
<td>495</td>
<td>0.8 (23%), 4.4 (77%)</td>
</tr>
<tr>
<td>2b</td>
<td>274 (0.89), 367 (0.38)</td>
<td>645</td>
<td>8</td>
<td>8</td>
<td>0.0001</td>
<td>0.0007</td>
<td>534</td>
<td>0.5 (20%), 3.8 (80%)</td>
</tr>
<tr>
<td>3a</td>
<td>273 (1.51), 330 (0.65)</td>
<td>504</td>
<td>79</td>
<td>216</td>
<td>0.011</td>
<td>0.036</td>
<td>458</td>
<td>0.9 (12%), 4.1 (88%)</td>
</tr>
<tr>
<td>3b</td>
<td>274 (9.55), 336 (0.57)</td>
<td>559</td>
<td>115</td>
<td>635</td>
<td>0.0055</td>
<td>0.024</td>
<td>488</td>
<td>3.2 (21%), 8.0 (79%)</td>
</tr>
</tbody>
</table>

* From air-equilibrated solutions. † From degassed solutions. ‡ Quinine sulfate in air-equilibrated 0.1 M H$_2$SO$_4$ used as the reference.

**Fig. 3** Experimental absorption (top) and emission profiles (bottom) of the complexes from dilute dichloromethane solutions (ca. 10$^{-5}$ M) at room temperature. The emission profiles were obtained by exciting the complexes to their lowest energy $\lambda_{abs}$. The same colour coding (inset) is used between the absorption and emission profiles.
The higher energy transitions in the region of 250–300 nm are calculated as being LC $\pi \rightarrow \pi^*$ in nature.

**Photochemical investigation**

The photochemical properties of the complexes as solutions in CD$_3$CN were assessed using $^1$H NMR spectroscopy with an irradiation wavelength of 365 nm, which corresponds to excitation into the MLCT/LLCT manifold. It was found that while the complexes with pyridyl-functionalised carbene ligands were photoactive, those containing pyrimidyl groups were photostable. This observation matches previously reported trends for rhenium complexes of carbene ligands.$^{14,15}$ The $^1$H NMR spectra of 2a during photolysis over 8 hours is shown in Fig. 5. Examination of signals associated with N6$^+$ shows that, after one hour, complex 2a has disappeared and two new species of low and high integration have been formed at 8.63 and 8.71 ppm, respectively. The signal at 8.63 has the same integration as another new signal at 9.44 ppm, which is characteristic of the H2 hydrogen in an imidazolium salt and suggests that these signals are associated with the formation of 1-pyridyl-3-mesitylimidazolium azide. It seems probable that the H2 hydrogen atom of the imidazolium salt arises from adventitious water present in the deuterated solvent. Over the course of the photolysis, the signals associated with the imidazolium salt continue to grow, while the intensity of the signal at 8.71 ppm increases for 3 or 4 hours, then decreases to zero by 7 hours. Close examination also shows that after 4 hours, a new species at 8.85 ppm begins to grow. Although the chemical shift of this new photoproduct is the same as for 2a, it is clearly a different compound as is evident by the disappearance of 2a after only one hour. The formation of unidentifiable decomposition products was observed from 8.87–8.98 ppm.

The rapid photolysis of 2a shown in Fig. 5 does not give an indication of what occurs in the early stages of the photoreaction. Therefore, the experiment was carefully repeated for the initial hour to see if any intermediates were formed that might shed some light on the reaction pathway. The results were illuminating, and are shown in Fig. 6. After one minute of photolysis, the imidazolium salt can just be observed as signals at 9.44 and 8.63 ppm, and continues to slowly form during the entire 60 minutes. The next products to form are first observed after 5 minutes photolysis at 8.81 and 8.70 ppm; the former growing very slowly and the latter at 8.71 growing to be the major product over the course of the hour.

Although the NMR photolysis of 2a shows the formation of at least five different products, their identities could not be properly determined, except for the easily recognisable imidazolium salt 1-pyridyl-3-mesitylimidazolium azide. Monitoring the photolysis by IR spectroscopy (Fig. 7), however, gave a better understanding of the processes occurring. The final trace shows the presence of two pairs of carbonyl peaks at 1938, 1866 cm$^{-1}$ and 1916, 1833 cm$^{-1}$, the latter pair being of much higher intensity. The bands at 1938, 1866 cm$^{-1}$, however, only began to form after 3 hours photolysis. Residual peaks of lower intensity are also observed at 2036 and 2017 cm$^{-1}$. From comparison to our earlier reports of similar complexes with halide ligands instead of azide, the IR profile of photolyzed 2a allows us to assign all of the photoproducts.$^{14}$ The signals at 1916 and 1833 cm$^{-1}$ are consistent with the formation of...
the neutral complex [Re(NHC)(CO)₂(CH₃CN)N₃] (I), while the smaller pair of carbonyl bands are assigned to [Re(NHC)(CO)₂(CH₃CN)₂]N₃ (II). Interestingly, and in contrast to the photochemistry of our previously reported rhenium NHC complexes, the neutral species (I) is the major product formed by photolysis, while the ionic complex (II) is only formed in very small amounts. Close examination of the band at 2036 cm⁻¹ shows that the intensity grows up to 3 hours and then decreases. This is consistent with the formation of the intermediate [Re(NHC)(CO)₂(CH₃CN)]N₃ (III), which is itself consumed in the formation of (II). Indeed, this band is located at the same wavenumber as the [Re(NHC)(CO)₃(CH₃CN)]Cl, where the NHC ligand contains a phenyl substituent instead of mesityl.

The absorption and emission profiles of 2a were measured prior to and after exhaustive photolysis, and are shown in Fig. 8. Significant changes have occurred in the absorption spectra after photolysis, with a loss of the defined MLCT/LLCT band and an increase in the absorbance above 400 nm. The emission spectra shows an initial band centred at 557 nm followed by a large drop in the emission intensity upon photolysis consistent with the photo decomposition of 2a, and the formation of a broad and structureless band that is slightly red-shifted. A red-shifted band of lowered emission intensity is in agreement with the proposed formation of [Re(NHC)(CO)₂(CD₃CN)N₃] (I), where the HOMO is raised due to the loss of one π-acidic carbonyl ligand.

The combination of the NMR, IR, absorption, and emission spectra allow the photochemistry of 2a in CH₃CN to be explained, and is shown in Scheme 2. After one minute photolysis, 2a begins to decompose to the imidazolium salt, seen as peaks in the 1H NMR spectra at 8.63 and 9.44 ppm. This observation warrant special mention, for although this side reaction proceeds steadily throughout the course of the photolysis, the formation of the imidazolium salt in this way does not seem consistent with a species that might form via the lowest MLCT excited state. As seen in Fig. 6, however, the imidazolium salt is clearly the first photoproduct formed, in a pathway independent of I-III. This is followed by the observance of (I) and (III) after only five minutes at 8.71 and 8.81 ppm, respectively, the former becoming the major species by 10 minutes of irradiation. As the photolysis proceeds (see Fig. 5 and 7), (I) continues to grow; the concentration of (III) remains low and...
after three hours the signal decreases to zero by ca. six hours irradiation. The decrease in the intensity of (III) is associated with the formation of a peak at 8.85 ppm due to complex (II), which is assumed to form via the substitution the carbonyl ligand by acetonitrile trans to the carbene.

The photoreactivity of 3a as a solution in CD$_3$CN was also assessed via $^1$H NMR spectroscopy, the resulting spectra summarised in Fig. 9. Compared to 2a, the reactivity of 3a upon irradiation was markedly slower, with no photoproducts being observed in the first two hours of the experiment. Examination of the N6$^+$ signal in the $^1$H NMR spectrum of 3a shows the presence of the N(2) and N(1) bound triazolate again in the ratio of 1:0.18, however this ratio increased over the 13 hours of irradiation. The exact amount of N(1) bound triazolate, however, could not be determined due the appearance of an unidentified species at 8.69 ppm that partially overlapped the signal. In a similar fashion to the photolysis of 2a, signals at 8.63 ppm due to an imidazolium salt formed after two hours and continued to increase in concentration during the experiment. It is not clear whether the counteranion of the imidazolium salt would be the triazolate or a triazolate-derived decomposition product. Monitoring the photolysis by IR spectroscopy reveals little change over 6 hours, with the exception of two small bands being formed at 1867 and 1845 cm$^{-1}$. These bands may be attributed to the formation of a dicarbonyl complex, although we are uncertain at present of its identity. Still it seems likely that one or more dicarbonyl species are formed, given the number of species observed in the $^1$H NMR. In an attempt to monitor the emission of 3a during photolysis, a CH$_3$CN solution of the complex (ca. 10$^{-5}$ M) was exposed to prolonged excitation inside the fluorimeter, but only a minor decrease in the emission intensity and no noticeable shift in the wavelength of emission, was observed.

**Conclusion**

The synthesis of the first rhenium(i) N-heterocyclic carbene complexes containing azide ligands (2a and 2b) has been performed, with the subsequent cycloaddition reaction with model alkyne dimethyl acetylenedicarboxylate affording triazolate ligated complexes 3a and 3b. The photophysical properties of the compounds have been investigated, and revealed a dependence on both the heterocycle bound to the imidazole (pyridyl vs. pyrimidyl) and the auxiliary ligand (azide vs. triazolate). The complexes display phosphorescent decay from triplet charge transfer (metal-to-ligand and ligand-to-ligand) excited states with markedly different $\lambda_{em}$ and decay lifetimes ($\tau$). The pyrimidyl group serves to lower the relative energy of the LUMO due to being more electron deficient compared with the pyridyl group,
while the electron donating nature of the azide raises the HOMO cf. the electron withdrawing triazolate. Consistent with previous reports, the complexes with pyridyl substituents on the N-heterocyclic carbene were found to be photoactive when excited into the charge transfer manifold. Although difficult to interpret compared to analogous complexes containing halide ligands, photolysis of 2a revealed three independent pathways: the photo-decomposition to the imidazolium salt; formation of the neutral [Re(NHC)(CO)3(CD3CN)N3] (I) as the major species; and formation of the tricarbonyl acetonitrile solvate [Re(NHC)(CO)2(CD3CN)]N3 (II), which then rapidly reacts to give the dicarbonyl species [Re(NHC)(CO)3(CD3CN)]N3 (II). The photochemistry of 3a was much slower than that of 2a, with significant amounts of starting material still present after 13 hours irradiation. Again, the formation of an imidazolium salt and a general increase in the concentration of the N(1) bound-isomer was observed, but the other photoproducts have not yet been identified, although dicarbonyl complexes seem likely. The results presented in this paper describe new types of easily accessible rhenium(i) N-heterocyclic carbene complexes that have very different photophysical/photochemical properties to those previously reported. An investigation of the scope of the 1,3-dipolar cycloaddition reaction using rhenium(i) N-heterocyclic carbene complexes of azide ligands is currently underway in our group.

Experimental section

General considerations

All reagents and solvents were purchased from Sigma Aldrich and used as received without further purification. All reactions were conducted under an atmosphere of N₂. All reactions and subsequent manipulations were performed in the dark where possible. Complexes 1a and 1b were prepared according to previously published procedures. Nuclear magnetic resonance spectra were recorded using a Bruker Avance 400 spectrometer (400.1 MHz for ¹H; 100 MHz for ¹³C) at 300 K. All the NMR spectra were calibrated to residual solvent signals. Infrared spectra were recorded using an attenuated total reflectance Perkin Elmer Spectrum 100 FT-IR with a diamond crystal. IR spectra were recorded from 4000 to 650 cm⁻¹. The intensities of the IR bands are reported as strong (s), medium (m), or weak (w), with broad (br) bands also specified. Elemental analyses were obtained at Curtin University using a Thermo Finnigan EA 1112 Series Flash.

Photophysical measurements

Absorption spectra were recorded at room temperature using a Cary 4000 UV/Vis spectrometer. Uncorrected steady state emission and excitation spectra were recorded on an Edinburgh FLS980-S2S2-stm spectrometer equipped with: (i) a temperature-monitored cuvette holder; (ii) 450 W Xenon arc lamp; (iii) double excitation and emission monochromators; (iv) a Peltier cooled Hamamatsu R928P photomultiplier tube (spectral range 200–870 nm). Emission and excitation spectra were corrected for source intensity (lamp and grating) and emission spectral response (detector and grating) by a calibration curve supplied with the instrument. To record the 77 K luminescence spectra and emission lifetimes, the samples were put in glass tubes (2 mm diameter) and inserted in a special quartz Dewar filled up with liquid nitrogen. According to the approach described by Demas and Crosby, luminescence quantum yields (Φₑ) were measured in optically dilute solutions (O.D. < 0.1 at excitation wavelength) obtained from absorption spectra on a wavelength scale [nm] and compared to the reference emitter by the following equation:

\[ \Phi_x = \Phi_r \left( \frac{A_x(\lambda) \cdot I_x(\lambda) \cdot n_x}{A_r(\lambda) \cdot I_r(\lambda) \cdot n_r} \right)^2 \cdot \frac{D_x}{D_r} \]

where \( A \) is the absorbance at the excitation wavelength (\( \lambda \)), \( I \) is the intensity of the excitation light at the excitation wavelength (\( \lambda \)), \( n \) is the refractive index of the solvent, \( D \) is the integrated intensity of the luminescence and \( \Phi \) is the quantum yield. The quantum yields of complexes were measured against an air-equilibrated 0.1 M H₂SO₄ solution of quinine sulfate (\( \Phi_r = 0.53 \)).

Emission lifetimes (\( \tau \)) were determined with the time correlated single photon counting technique (TCSPC) with the same Edinburgh FLSP980-S2S2-stm spectrometer using either a pulsed picosecond LED (EPLED/EPL 377 nm, FHWM < 800 ps) or a microsecond flashlamp. The goodness of fit was assessed by minimizing the reduced \( \chi^2 \) function and by visual inspection of the weighted residuals. The solvents used for the preparation of the solutions for the photophysical investigations were of LR grade and the water was deionised. Experimental uncertainties are estimated to be ±8% for lifetime determinations, ±20% for quantum yields, ±2 nm and ±5 nm for absorption and emission maxima, respectively.

Monitored photolysis

Monitored photolysis experiments were carried out using an Edinburgh FLSP980-S2S2-stm spectrometer. The reagents were added to acetonitrile in order to reach a final concentration of ca. 10⁻⁵ M (~3 mL volume) in a quartz cuvette. The unstirred solutions were irradiated at \( \lambda_{ex} = 350 \text{ nm} \) with an excitation slit set to 10 nm and emission slit to 2.5 nm. A consecutive number of spectra, with no delay between replicates, were run per solution at a speed of 240 nm min⁻¹ with the detector set to acquire a signal between 400 and 800 nm so that the complexes underwent exhaustive photolysis. Under these experimental conditions, the solution was constantly irradiated for a period of approximately 3 hours. The spectra were recorded uncorrected for the detector response.

Lamp photolysis

Lamp photolysis experiments were carried using a UVP Blak-Ray B-100AP High Intensity UV lamp with a 100 W bulb at a single wavelength output of 365 nm. The experiments were performed in the darkness. The reaction vessel consisted of a either a glass NMR tube containing deuterated solution of the complex.
Numbering of complexes

**fac-[Re(CO)3(NHCpyridyl)N3]** (2a). Silver trflate (70 mg, 0.27 mmol) was added to a solution of 1a (103 mg, 0.18 mmol) in acetone (15 mL) and the resulting solution was heated at reflux for 3 h. The mixture was cooled, and then filtered under N2 via cannula filtration into another Schlenk flask. Sodium azide (118 mg, 1.81 mmol) was added and the mixture was stirred at room temperature overnight. The resulting solution was decanted and the solvent removed under reduced pressure at room temperature. The resulting residue was dissolved in dichloromethane and filtered through a plug of acidic alumina (BII), eluting with dichloromethane (~200 mL) and then dried under reduced pressure. The residue was dissolved in acetone (2 mL) and filtered through a plug of Celite into stirring pentane (20 mL). The resulting solid was collected, washed with pentane (4 x 3 mL), and dried to afford the product as a yellow solid. Yield: 39.3 mg, 68%. Anal. calc. for C19H16N7O3Re: C 43.29, H 2.73, N 11.01. FT-IR (ATR) νmax/cm⁻¹: 198.0 (CO), 195.1 (CO), 193.2 (CO), 191.2 (C2), 162.0 (C4' or C6'), 159.6 (C2'), 140.7 (mesityl quat. C), 136.7 (mesityl quat. C), 130.2 (mesityl CH), 129.9 (mesityl CH), 128.7 (mesityl CH), 119.4 (C4), 21.2 (mesityl CH), 17.8 (mesityl CH3), 17.6 (mesityl CH3), 17.6 (mesityl CH3). Crystals of 2b suitable for X-ray diffraction studies were grown by layering pentane on top of a dichloromethane solution of the complex and sealing the vial for several days.

**fac-[Re(CO)3(NHCpyridyl)(triazolate)]** (3a). Dimethyl acetylene-diacarbonylate (19 µL, 0.16 mmol) was added to a solution of 2a (26 mg, 0.045 mmol) in acetone (5 mL), and stirred at room temperature overnight. The resulting precipitate was collected, washed with pentane (4 x 2 mL), diethyl ether (3 mL), and dried to afford the product as an off-white solid. Yield: 17 mg, 52%. Anal. calc. for C26H25N6O4Re·0.15CH2Cl2: C 43.00, H 3.22, N 11.51; found: C 43.29, H 2.73, N 11.01. FT-IR (ATR) νmax/cm⁻¹: 1923 s (CO), 1905 s (CO), 1740 m, 1721 s, 1487 m, 1444 w, 1321 m. Major isomer: 1H-NMR (DMSO-d6): 8.86 (1H, d, JHH = 5.5 Hz, H6'), 8.79 (1H, d, JHH = 2.2 Hz, H4), 8.35–8.51 (2 x 1H, 2 x m, H3' and H4'), 7.78 (1H, d, JHH = 2.2 Hz, H5), 7.52 (1H, d, JHH = 6.7 Hz, JHH = 5.5 Hz, JHH = 2.2 Hz, H5'), 7.16 (1H, s, mesityl CH), 7.04 (1H, s, mesityl CH), 3.70 (6H, s, triazole CH3), 2.33 (3H, s, mesityl CH3), 2.13 (3H, s, mesityl CH3), 1.56 (3H, s, mesityl CH3). 13C-NMR δ/ppm (DMSO-d6): 198.98 (CO), 196.18 (CO), 193.98 (C2), 192.19 (CO), 154.3 (C6'), 140.6 (mesityl quat. C), 143.0 (C4'), 136.7 (mesitylquat. CH), 136.4 (mesitylquat. CH), 135.6 (mesityl quat. CH), 130.2 (mesityl CH), 129.9 (mesityl CH), 129.5 (C5), 125.1 (C5'), 118.8 (C4), 113.8 (C3'), 21.2 (mesityl CH3), 17.8 (mesityl CH3), 17.6 (mesityl CH3), C2' was not observed. Crystals of 3a suitable for X-ray diffraction studies were grown by the diffusion of vapours between pentane and a methanol/acetone/dichloromethane solution of the complex.

**fac-[Re(CO)3(NHCpyridyl)]** (2b). Silver trflate (60 mg, 0.23 mmol) was added to a solution of 1a (63.5 mg, 0.10 mmol) in acetone (10 mL) and the resulting solution was heated at reflux for 3.5 h. The mixture was cooled, and then filtered under N2 via cannula filtration into another Schlenk flask. Sodium azide (72 mg, 1.11 mmol) was added and the mixture was stirred at room temperature overnight then at reflux for 1.5 h. The resulting solution was decanted and the solvent removed under reduced pressure at room temperature. The resulting residue was dissolved in dichloromethane and filtered through a plug of acidic alumina (BII), eluting with dichloromethane (~200 mL) and then dried under reduced pressure. The residue was dissolved in acetone (2 mL) and filtered through a plug of Celite into stirring pentane (20 mL). The resulting solid was collected, washed with pentane (4 x 3 mL), and dried to afford the product as a yellow solid. Yield: 39.3 mg, 68%. Anal. calc. for C19H16N7O3Re·0.30C3H6O: C 39.6, H 2.97, N 16.24; found: C 40.0, H 2.24, N 15.71. FT-IR (ATR) νmax/cm⁻¹: 2056 s (N3), 2007 s (CO), 1889 br s (CO), 1712 w, 1595 s, 1566 w, 1468 s, 1378 m. 1H-NMR δ/ppm (acetone-d6): 9.30 (1H, d, JHH = 5.6 Hz, H5'), 9.19 (1H, d, JHH = 4.8 Hz, JHH = 2.2 Hz, H4' or H6'), 8.36 (1H, d, JHH = 2.2 Hz, H4), 7.76 (1H, d, JHH = 5.6 Hz, JHH = 4.8 Hz, H5'), 7.60 (1H, d, JHH = 2.2 Hz, H5), 7.12 (2 x 1H, 2 x s, 2 x mesityl CH), 2.36 (3H, s, mesityl CH), 2.18 (3H, s, mesityl CH), 2.15 (3H, s, mesityl CH), 2.15 (3H, s, mesityl CH3), 1.56 (3H, s, mesityl CH3). FT-IR (ATR) νmax/cm⁻¹: 1951 s (CO), 1930 s (CO), 1904 (C4), 1630 (C3' or C6'), 159.6 (C2'), 140.7 (mesityl quat. C), 136.7 (mesitylquat. C), 130.2 (mesityl CH), 129.9 (mesityl CH), 128.7 (mesityl CH), 119.4 (C4), 21.2 (mesityl CH), 17.8 (mesityl CH3), 17.6 (mesityl CH3). Crystals of 2b suitable for X-ray diffraction studies were grown by layering pentane on top of a dichloromethane solution of the complex and sealing the vial for several days.
was too low to observe the carbon resonances in the $^{13}$C NMR spectrum.

$^{13}$C-NMR (DMSO-$_d$$_6$): 197.3 (CO), 192.5 (CO), 190.6 (C$_2$), 187.3 (CO), 186.4 (C$_2$), 163.1 (C$_4$’ or C$_6$’), 161.3 (C$_4$’ or C$_6$’), 139.0 (triazolate quat. C), 135.2 (mesityl C), 139.1 (mesityl quat. C), 118.7 (C$^4$), 51.7 (triazolate CH$_3$), 20.6 (mesityl CH$_3$), 17.2 (mesityl CH$_3$), 16.0 (mesityl CH$_3$). Minor isomer: $^1$H-NMR $^\delta$/ppm (DMSO-$_d$$_6$): 9.21 (1H, d, $^3$J$_{HH}$ = 2.1 Hz, H$_4$’), 9.14 (1H, dd, $^3$J$_{HH}$ = 5.6 Hz, $^3$J$_{HH}$ = 2.1 Hz, H$_4$’ or H$_6$’), 8.57 (1H, d, $^3$J$_{HH}$ = 2.2 Hz, H$_4$), 7.77 (1H, d, $^3$J$_{HH}$ = 2.2 Hz, H$_5$’), 7.61 (1H, m, H$^5$), 7.15 (1H, s, mesityl CH$_3$), 7.01 (1H, s, mesityl CH), 3.79 (3H, s, mesityl CH$_3$), 3.72 (3H, s, triazolate CH$_3$), 2.13 (3H, s, mesityl CH$_3$), 2.12 (3H, s, mesityl CH$_3$), 1.24 (3H, s, mesityl CH$_3$). The concentration of the N1-linked triazolate complex (minor isomer) in the sample was too low to observe the carbon resonances in the $^{13}$C NMR spectrum.

X-ray diffraction analysis. Crystallographic data for the structures were collected at 100(2) K on an Oxford Diffraction Xcalibur fitted with Mo K$_{\alpha}$ radiation $\lambda$ = 0.71073 Å Following analytical absorption corrections and solution by direct methods, the structures were refined against $F^2$ with full-matrix least-squares using the program SHELXL-2014. Anisotropic displacement parameters were employed for the non-hydrogen atoms. All hydrogen atoms were added at calculated positions and refined by use of a riding model with isotropic displacement parameters based on those of the parent atom.

Crystal data

2a. Formula: C$_{26}$H$_{23}$N$_7$O$_7$Re, $M$ = 717.70. Monoclinic, space group $F21/c$, $a$ = 11.3508(2), $b$ = 14.0357(2), $c$ = 16.9573(2) Å, $\beta$ = 95.1090(10)°, $V$ = 2597.18(7) Å$^3$, $Z$ = 4, $D_{calc}$ = 1.806 Mg m$^{-3}$, $\mu$ = 4.655 mm$^{-1}$, $2\theta_{max}$ = 88°, crystal size = 0.16 × 0.10 × 0.04 mm$^3$. Reflections collected = 73048, unique = 10783, $R_{int}$ = 0.0450. Data/restraints/parameters = 10783/30/361, GeoF = 1.052. Final $R$ indices ($I > 2\sigma(I)$): $R_I$ = 0.0177, $wR_I$ = 0.0573. $R$ indices (all data): $R_I$ = 0.0362, $wR_I$ = 0.0607. $\Delta P_{max/min}$ = 1.444, −0.785 e Å$^{-3}$. CCDC 1431848.

Computational calculations

Time-dependent density functional theory (TD-DFT) calculations were performed with GAUSSIAN 09 in order to calculate the absorption spectra for synthesized complexes. Prior to these calculations, the structures were relaxed at the B3LYP level of theory in the presence of an implicit solvent (dichloromethane). The low-lying singlet–triplet excitation energies were calculated at the same level of theory, and the spectra were reproduced as the superposition of Gaussian functions with heights proportional to calculated intensities and a variance of 11 nm.

Acknowledgements

This work is financially supported by the Australian Research Council (FT130100033 to MM and FT130100463 to PR).
The authors acknowledge the facilities, and the scientific and technical assistance of the Australian Microscopy & Microanalysis Research Facility at the Centre for Microscopy, Characterisation & Analysis, The University of Western Australia, a facility funded by the University, State and Commonwealth Governments. This work was supported by resources provided by the National Computational Infrastructure (NCI), which is supported by the Australian Government.

References