

Light Activated Electrochemistry: Light Intensity and pH Dependence on Electrochemical Performance of Anthraquinone Derivatized Silicon

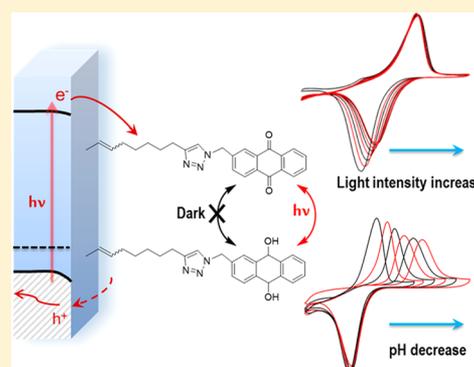
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S Supporting Information

ABSTRACT: We seek to understand how the thermodynamics and kinetics of anthraquinone-containing self-assembled monolayer on silicon electrodes are affected by two key experimental variables: the intensity of the light assisting the anthraquinone/anthrahydroquinone redox process and the local solution environment. The substrates are chemically passivated poorly doped p-type silicon electrodes. The study presents a strategy for the selective modulation of either the anodic or the cathodic process occurring at the interface. Cyclic voltammetry studies showed that unlike for a proton-coupled electron transfer process performed at metallic electrodes, for the redox reaction of the anthraquinone unit on a silicon electrode it becomes possible to (i) selectively facilitate only the oxidation process by increasing the electrolyte pH or (ii) at a given pH value to increase the illumination intensity to anodically shift the onset of the reduction step only but leave the oxidation process thermodynamic unchanged. A model concerning the proton coupled electron transfer mechanism was proposed, where the electron transfer is the rate-determining step for the anthraquinone reduction while a deprotonation step is the rate-determining event for the anthrahydroquinone oxidation on poorly doped illuminated p-type silicon.



1. INTRODUCTION

Fundamental studies into semiconducting electrodes were pioneered in the 1970s and 1980s by researchers including Bard, Wrighton, and Lewis.^{1–8} These early works formed the basis for the use of silicon for the light-assisted reduction of CO₂,⁹ water splitting,¹⁰ and hydrogen generation.^{2,7,11} These technologies all rely on silicon being the photoresponsive material in which, upon the adsorption of light, electron–hole pairs are generated both in the field-free bulk and depletion region near the silicon/electrolyte interface (space charge). Carriers in the depletion layer can be separated by the electric field such that minority carriers are directed to the surface while the majority carriers are directed toward the bulk.¹² Based on this principle, the obvious possibility of a two-dimensional confinement for the photoresponse led to the development of light addressable potentiometric sensors (LAPS) for biochemical monitoring¹³ and multianalyte detection.^{14–16} With regards to the former example, LAPS, the typical device consists of an electrolyte/insulator/semiconductor structure. An alternating photocurrent response is measured when the LAPS device is irradiated with a frequency-modulated light pointer at a specific position. The presence of charged species adjacent to the insulator surface influences the local characteristics of the space charge layer and hence the magnitude and phase of the alternating current. Thus, by measuring the amplitude or the

phase of the photocurrent, a LAPS devices can realize detection of specific ions and biological species with the two-dimensional resolution down to $17 \pm 3 \mu\text{m}$.¹⁷

The main limitation of LAPS is that only potentiometric sensing can be performed. This is because the insulating layer adds a significant kinetics barrier to electron transfer from and to the electrolyte. To broaden the scope of such an “addressing” concept in electroanalytical chemistry would require a method to allow faradaic electrochemistry to be performed specifically only at the site of illumination. This is what was achieved with the light activated electrochemistry concept we reported recently.¹⁸ With light activated electrochemistry, faradaic reactions can be performed wherever and whenever you want, with microscale resolution on an unstructured macroscopic monolithic silicon electrode simply by illuminating a discrete spot on that surface. To achieve light-activated electrochemistry, three conditions must be fulfilled. First the semiconducting electrode must be biased in depletion, second the electrode must be modified such that it does not undergo photooxidation during electrochemistry but with a layer thin enough to allow efficient tunneling, and finally there

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must be sufficient electronic coupling with the redox active species as found when it is covalently bound to the surface-modifying layer. In our preliminary study,¹⁸ we showed that these conditions could be fulfilled on poorly doped n-type “oxide-free” silicon electrodes that were modified with a self-assembled monolayer of 1,8-nonadiyne, to which a ferrocene unit was subsequently attached. This idea was shown to be applicable to writing electrochemical information to a surface, via the electrodeposition of polypyrrole, and to read electrochemical information from a discrete area of a macroscopic surface in terms of DNA hybridization sensing.

The conditions discussed above exist because the core principle of light-activated electrochemistry is that there are insufficient carriers for the electrochemistry to proceed in the dark, and irradiation causes photoexcitation of electrons (for a p-type silicon) from the valence band to the conduction band. There is a concomitant increase in the number of the minority carriers such that charge transfer can then proceed at an appreciable rate, and faradaic processes are observed. To further understand the scope of light-activated electrochemistry, however, it is important to examine the impact of (i) the light intensity on the faradaic process and (ii) with redox reactions that are environmentally sensitive, such as for proton-coupled electron transfer reactions, the interplay between photoeffects, and electrolyte composition on the electrochemistry. These effects will be investigated using nonadiyne modified p-type electrodes functionalized with anthraquinone via Cu(I)-catalyzed alkyne–azide cycloaddition reaction (“click” or CuAAC).^{19–21} In this way, if a poorly doped p-type silicon electrode is utilized, it will only be able to sustain electron transfer to the anthraquinone at an appreciable rate when illuminated.

As a typical proton-coupled electron transfer process, the redox reaction of quinones has been studied extensively, both experimentally and theoretically.^{22–28} Laviron²² and Finklea^{23,24} proposed the nine-number square scheme on proton-coupled electron transfer reactions to account for a number of experimental observations. The stepwise mechanism of proton-coupled electron transfer is largely accepted. However, there are some recent experimental results obtained on proton coupled electron transfer reactions that do not fit with the stepwise mechanism of the *square* scheme;^{25,26} hence the possibility of concerted electron and proton transfer was introduced.^{25,27} The exact proton-coupled electron transfer mechanism for anthraquinone redox process is therefore still a source of debate,²⁷ however, there is general agreement on both thermodynamics and the kinetics being pH dependent.^{23,28–31} The redox reaction of quinone and phenols derivatives do follow a shift in potential according to the Nernst equation,^{28,29,32} and the magnitude of the redox potential shift versus pH can be used to extract system parameters such as the number of electrons involved in the redox reaction. For quinones two electrons and two protons were involved in the redox reaction and slopes closed to -59.2 mV pH^{-1} for equilibrium redox potential versus pH plots were observed.^{28,31}

Covalent attachment of on electrodes to form self-assembled monolayers allows controlling the local environment at the site of the redox reaction. There have been several studies on anthraquinone-terminated self-assembled monolayers on gold, glass carbon, and nickel electrodes.^{29,30,33–35} However, the electrochemistry for anthraquinone-terminated self-assembled monolayers at semiconducting silicon in aqueous environment

is still unexplored. Therefore, the purpose of this work is to modify anthraquinone-terminated self-assembled monolayers on silicon substrate and to determine the effects of pH and light intensity on the thermodynamics and kinetics of anthraquinone redox reactions. The anodic and cathodic peak variations of anthraquinone-terminated silicon at different pH values as a function of light intensity were investigated by cyclic voltammetry. The rate constants for the electron transfer as well as the changes in space-charge characteristics with variation in illumination intensity were studied by electrochemical impedance spectroscopy.

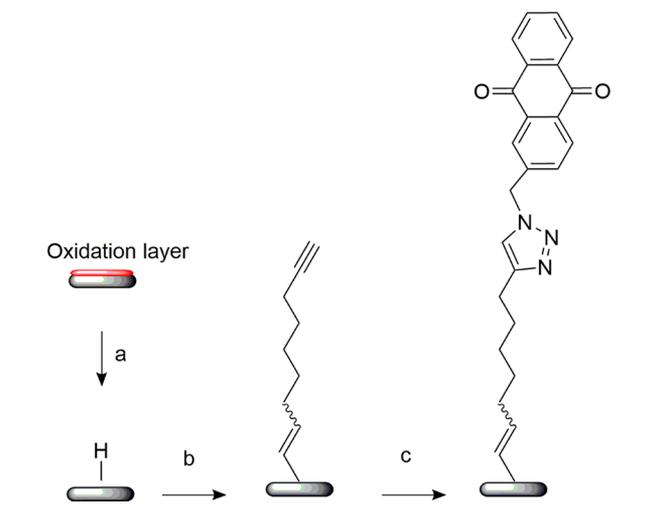
2. EXPERIMENTAL METHODS

2.1. Chemicals and Materials. Unless otherwise specified, all chemicals are of analytical grade and used as received. Solvents used for the silicon cleaning steps were redistilled prior to use. Hydrogen peroxide (Sigma-Aldrich, 30 wt % in water), hydrofluoric acid (Riedel-Haën, 48 wt % in water), and sulfuric acid (J.T. Baker, 96%) were of semiconductor grade. 1,8-Nonadiyne (Sigma-Aldrich, 98%) was redistilled from sodium borohydride (Sigma-Aldrich, $\geq 99\%$) under reduced pressure (65–70 °C, 25–30 Torr) and stored under a dry argon atmosphere prior to use. The redox species 2-(azidomethyl)anthracene-9,10-dione was synthesized in one step from 2-(bromomethyl)anthracene-9,10-dione. Silicon substrates were prime grade, single-side polished, 100-oriented ($\langle 100 \rangle \pm 0.9^\circ$), p type (boron), with nominal resistivity of 0.007–0.009 and 10–20 $\Omega \text{ cm}$ and are referred to as highly and poorly doped, respectively. The poorly doped silicon were $500 \pm 25 \mu\text{m}$ thick and obtained from Virginia Semiconductors, Inc. (Fredericksburg, VA). The highly doped silicon were $525 \pm 25 \mu\text{m}$ thick and were purchased from Siltronix, S.A.S. (Archamps, France).

2.2. X-ray Photoelectron Spectroscopy. X-ray photoelectron spectra were obtained on an ESCALAB 220iXL spectrometer fitted with a monochromatic Al $K\alpha$ source (1486.6 eV), a hemispherical analyzer, and a six multichannel detector. The data were collected in normal emission at room temperature and a pressure below 10^{-9} mbar. The resolution of the spectrometer was $\sim 0.6 \text{ eV}$ as measured from the Ag $3d_{5/2}$ signal full width at half-maximum. All binding energies are corrected by applying a rigid shift to bring the C 1s signal to 285.0 eV. The Thermo Avantage software was used for the analysis of XPS data. The atomic compositions were corrected for atomic sensitivities and measured from high-resolution scans. Atomic sensitivity factors are often instrument-sensitive, and the atomic sensitivities obtained from the Avance software interfaced with the spectrometer were 0.817 for Si 2p, 1.000 for C 1s, and 1.800 for N 1s.

2.3. Surface Modification Procedures. **2.3.1. Passivation of Silicon(100) with Monolayers of 1,8-Nonadiyne.** The assembly of monolayers of 1,8-nonadiyne was performed on silicon following a previously reported procedure.^{20,36} As shown in [Scheme 1](#), a silicon wafer (ca. $10 \times 30 \text{ mm}$) was rinsed with dichloromethane, dried under a stream of argon, and then cleaned in hot piranha solution (100 °C, a 1:3 (v/v) mixture of 30% aqueous hydrogen peroxide to concentrated sulfuric acid for 30 min. (**Caution: piranha solution is dangerous when hot and explosive when mixing with organic compounds.**) The sample was subsequently rinsed with water and etched with 2.5% aqueous hydrogen fluoride solution for 90 s, before being transferred to a neat 1,8-nonadiyne sample which had been degassed through a

Scheme 1. Surface Modification Processes: (a) Single-Side Polished Silicon(100) Wafers Were Immersed in 2.5% HF for 90 s to Remove the Silicon Oxidation Layer; (b) Obtained Hydride-Terminated Surface Was Reacted with a Neat Sample of 1,8-Nonadiyne via Hydrosilylation Process; (c) Copper-Catalyzed Azide–Alkyne Cycloaddition (CuAAC) Reaction of an Azide-Tagged Anthraquinone Derivative on Acetylene-Terminated Silicon(100) Surface



minimum of five freezing–pump–thawing cycles. The sample was kept under an argon atmosphere while the reaction flask was immersed for a 3 h period in an oil bath set to 165 °C. The reaction vessel was allowed to cool to room temperature, and the passivated silicon wafer was rinsed several times with dichloromethane, ethanol, immersed in dichloromethane at 4 °C for 2 h, blown dry under argon, and stored under argon for further processing or analysis.

2.3.2. “CuAAC” Attachment of Redox-Active 2-(Azidomethyl)anthracene-9,10-dione. Anthraquinone was attached on alkyne-functionalized silicon(100) samples through the CuAAC reactions (Scheme 1).^{19,37} The passivated silicon sample was placed in a dimethyl sulfoxide solution containing 2 mM 2-(azidomethyl)anthracene-9,10-dione, 9.76 mM copper(I) bromide, and 20.6 mM tetramethylethylenediamine. Reactions were carried out at room temperature for 30 min without excluding air from the reaction environment. The prepared silicon samples were rinsed consecutively with copious amounts of water, ethanol, and dichloromethane, blown dry under argon, and stored under argon before being analyzed.

2.4. Electrochemical Measurements. All electrochemical experiments were performed in a PTFE three-electrode single-compartment cell with the modified silicon electrode as the working electrode, a platinum mesh as the counter electrode, and Ag/AgCl in 1 M KCl as the reference electrode. All potentials are reported versus the reference electrode. A rectilinear cross-section Viton gasket defined the geometric area of the working electrode to 26 mm². The bulk of the backside of the silicon sample was exposed with emery paper and rubbed with gallium indium eutectic. A copper electrode was pressed against the sample backside, and it served as the ohmic contact. The cell was enclosed in a grounded Faraday cage during all measurements. The cyclic voltammetry were performed using a BAS 100B electrochemical analyzer (Bioanalytical Systems, Inc., West Lafayette, IN). The apparent

rate constant for electron transfer (k_{et}') was calculated from the simplified circuit³⁸ of electrochemical impedance spectroscopy; the values were in reasonable agreement with the values from Laviron method (Table S1, Supporting Information).³⁹ Flat band potential values were calculated from the extrapolation to $C_{sc}^{-2} = 0$ of Mott–Schottky ($M-S$) plots (C_{sc}^{-2} vs E). Electrochemical impedance spectroscopy under different light intensity was recorded using Solartron model SI 1287 with a Solartron impedance/gain-phase analyzer. The half-wave potential ($E_{1/2}$) was set as dc bias potential, and the ac amplitude was 10 mV. Experiments were performed in Britton Robinson (B&R) buffer⁴⁰ with a pH range 5.06–10.09. All the buffer solutions were degassed prior to experiments by bubbling argon through them for 20 min.

ACE Light Source (SCHOTT AG, Mainz, Germany) was employed to perform the top-side (i.e., electrolyte side) illumination for electrochemical experiments. The light intensities were determined using a 1918-R optical power meter (Newport, Inc., Bozeman, MT). All through the present work, the illumination was performed from the solution side of the electrode surface, henceforth referred to as top-side illumination.

3. RESULTS

3.1. Spectroscopic Characterization of Anthraquinone Assemblies on Silicon. The chemical passivation of the silicon surface has been extensively characterized,^{20,36,41,42} and representative XPS spectra for the alkyne-terminated poorly doped p-type silicon(100) surface are shown in Figure S1. In addition to the bands of Si, signals are observed for C and O. In high-resolution Si 2p spectra, the spin–orbit split peaks for Si 2p_{3/2} and Si 2p_{1/2} are shown at 99.5 and 100.1 eV, respectively. Three C 1s features from C 1s core level spectra at 283.77 (C–Si), 285.0 (C–C), and 286.4 eV (C–O–H) are present on 1,8-nonadiyne passivated surfaces as discussed previously.^{21,36,42}

The further modification of the alkyne-terminated silicon electrode by attachment of anthraquinone derivative (Scheme 1) was spectroscopically monitored by XPS (Figure S2). The utilization of CuAAC reactions on alkyne-terminated silicon surface for grafting azide-tagged molecules has been detailed extensively in previous studies.^{14–16} Concurring with expectations, the high-resolution N 1s XPS spectra (Figure S2a) for anthraquinone modified silicon are fitted to two distinct peaks having binding energies of 401.7 and 400.3 eV and a 1:2 ratio of the integrated area, which are corresponding to the triazole N–N=N and C–N–N, respectively.^{19,20,37} The presence of a triazole group is evidence for the success for alkyne–azide cycloaddition reaction. The success of the reaction was further verified by high-resolution C 1s XPS narrow scans (Figure S2c). The low-energy peak centered at 283.8 was attributed to C–Si, while the C 1s peaks with binding energy of 285.0 and 284.5 eV were ascribed to carbon-bonded carbon atoms in aliphatic and aromatic configuration, respectively. The two high-energy peaks centered at 288.0 eV (C=O for carbonyl group) and 286.5 eV (C–N) are also consistent with the anthraquinone modification on silicon.^{42–45} Importantly, the lack of XPS emissions in the energy range 101–105 eV indicates there is no appreciable substrate oxidation upon the CuAAC reaction.^{46,47}

3.2. Light and Dark Redox Behavior of Anthraquinone Derivatized p-Type Silicon Electrodes. The redox response of the anthraquinone-functionalized p-type silicon was first studied using cyclic voltammetry (Figure 1). As evident from

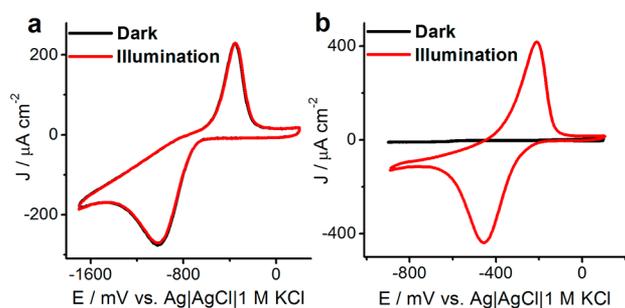


Figure 1. Cyclic voltammograms for anthraquinone-functionalized (a) highly ($0.007\text{--}0.009\ \Omega\ \text{cm}$ resistivity) and (b) poorly doped ($10\text{--}20\ \Omega\ \text{cm}$ resistivity) p-type silicon(100) in B&R buffer (pH = 10.06) in dark or under illumination. Scan rate was 2 V/s.

the cyclic voltammograms in aqueous electrolytes, the protonation of the radical anion collapses the anthraquinone reduction reactions to a single two-electron process.³³ For highly doped p-type silicon there is no appreciable difference between voltammograms in the dark or under illumination (Figure 1a). Conversely, for poorly doped p-type silicon electrodes in the dark there is no evidence within the same potential window of any faradaic process, which accords well with the previous findings for poorly doped p-type silicon (Figure 1b).^{48,49} However, illumination of the electrode with light of energy greater than the band gap leads to dramatic changes in the voltammetry. In the dark, electrons are in depletion at the bias required for the reduction of anthraquinone to hydroquinone. Illumination generates electron–hole pairs and significantly increases the number of electrons in conduction band, which will be further driven to the silicon/electrolyte interface.^{12,50} As a consequence, there are sufficient carriers for electron transfer to proceed at anthraquinone-terminated silicon/electrolyte interface; hence, faradaic electrochemistry is observed. Compared with cyclic voltammograms of anthraquinone on metal electrodes^{29,51,52} the anodic and cathodic peak potentials (E_{pa} and E_{pc}) for anthraquinone on poorly doped p-type silicon appear at more positive potentials. For example, the $E_{1/2}$ for anthraquinone derivative on gold electrode was found at pH 10 approximately $-550\ \text{mV}$ vs Ag|AgCl 3 M KCl,²⁹ while on poorly doped p-type silicon, light can drive the anthraquinone electrochemical reaction at $\sim 220\ \text{mV}$ positive than the $E_{1/2}$ of the couple on gold electrode at the same pH value. From analysis of the area under the anodic wave, and considering a two-electron process, the surface coverage of anthraquinone moieties was estimated to be $(1.49 \pm 0.19) \times 10^{-10}\ \text{mol}\ \text{cm}^{-2}$, which is lower than the reported value (ca. $5 \times 10^{-10}\ \text{mol}\ \text{cm}^{-2}$) for fully covered monolayers of hydroquinone on gold surface.⁵³

3.3. Effects of Electrolyte pH on the Voltammetry of Anthraquinone Derivatized Poorly Doped p-Type Silicon. A representative set of cyclic voltammograms for anthraquinone modified p-type silicon were recorded in buffered electrolytes of pH values between 5.06 and 10.09 (Figure 2a and Figure S3a). Our data showed on anthraquinone modified highly doped p-type silicon that $E_{1/2}$ shifts with a slope of $-58.5\ \text{mV}\ \text{pH}^{-1}$ (Figure S3b). In contrast, the $E_{1/2}$ on poorly doped p-type silicon shifts to more negative values by $-44.0\ \text{mV}\ \text{pH}^{-1}$. The enlarged separation between the oxidation and reduction peaks with a decrease of pH implies the possible electrode kinetic changes. A quantification of the apparent rate constant for the electron transfer, k_{et}' , was

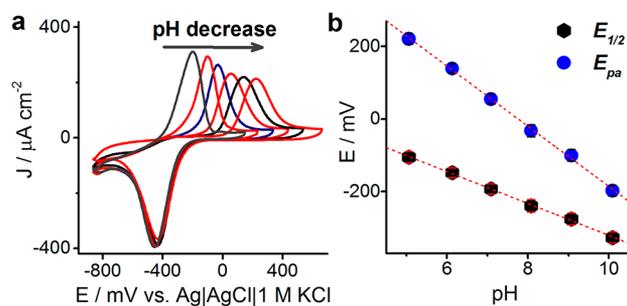


Figure 2. (a) Cyclic voltammetry of anthraquinone-modified poorly doped, p-type silicon at different pH values under illumination (light intensity = $94.2\ \text{mW}\ \text{cm}^{-2}$, with scan rate 2 V/s). (b) Dependence of half-wave potential ($E_{1/2}$) and anodic peak potential (E_{pa}) values with pH. Dashed lines in (b) are plots of the fitted linear equations, $E_{1/2} = 119.2 - 44.0\text{pH}$ and $E_{pa} = 644.3 - 83.0\text{pH}$.

measured by electrochemical impedance spectroscopy (EIS). The k_{et}' on anthraquinone-terminated poorly doped p-type silicon was found to decrease from $4.4\ \text{s}^{-1}$ at pH 10.09 to $0.02\ \text{s}^{-1}$ at pH 5.06, while the k_{et}' on anthraquinone-terminated highly doped p-type silicon decreased from 1.4 to $0.01\ \text{s}^{-1}$ (Table S1 and Figure S5a). The trend line for logarithmic k_{et}' versus pH was consistent with the previous studies.^{29,54,55} Uniquely, the influence of pH on E_{pa} and E_{pc} for anthraquinone modified silicon is quite different. The dependency of E_{pc} on the pH was linear with a very small slope of ca. $-5.2\ \text{mV}\ \text{pH}^{-1}$, about 15 times smaller than dependency of E_{pa} on the pH, which was $-83.1\ \text{mV}\ \text{pH}^{-1}$. In previous studies, the influence of pH on the thermodynamics of anthraquinone-terminated self-assembled monolayers on gold electrodes followed the Nernst equation,^{29,53–55} with a shift in apparent formal potential of $-59.2\ \text{mV}\ \text{pH}^{-1}$ between pH 1 and 13.^{31,53,55} The electrochemical reaction was shown to be a two-electron and two-proton reaction. Our data at anthraquinone modified poorly doped p-type silicon electrodes showed a pH dependence of $E_{1/2}$ is significantly smaller. This is because of the near negligible pH dependence of the cathodic ($-5.2\ \text{mV}\ \text{pH}^{-1}$) peak even though the pH dependence of the anodic peak is much greater than expected ($-83.1\ \text{mV}\ \text{pH}^{-1}$).

3.4. Light Intensity Dependence on the Voltammetry of Anthraquinone Derivatized Poorly Doped p-Type Silicon. To further explore the photoeffects on the redox reaction of anthraquinone moieties on poorly doped silicon, cyclic voltammetry was performed under variable light intensities (10.3, 16.7, 25.8, 38.2, 52.0, 69.6, 86.7, and $94.2\ \text{mW}\ \text{cm}^{-2}$) in B&R buffer at a pH of 10.1 (Figure 3). Typical voltametric curves of light-driven anthraquinone redox processes are shown in Figure 3a. The increased light intensity drives the $E_{1/2}$ to more positive potential, with E_{pa} around $-200 \pm 4\ \text{mV}$ showing no obvious response to light intensity, while E_{pc} moved toward positive potential with the increase of light intensity. Figure 3b shows light intensity dependence of E_{pc} . The difference in E_{pc} was as much as $88.1 \pm 5.8\ \text{mV}$ as the light intensity was changed from 10.3 to $94.2\ \text{mW}\ \text{cm}^{-2}$.

The electrode kinetics as a function of the light intensity was also studied, and it is found that the k_{et}' for the electron transfer process increased with increasing light intensities (Table S1 and Figure S5b). Specifically, the experimental obtained k_{et}' increased from $0.5\ \text{s}^{-1}$ at a light intensity of $10.3\ \text{mW}\ \text{cm}^{-2}$ to $4.4\ \text{s}^{-1}$ at a light intensity of $94\ \text{mW}\ \text{cm}^{-2}$.

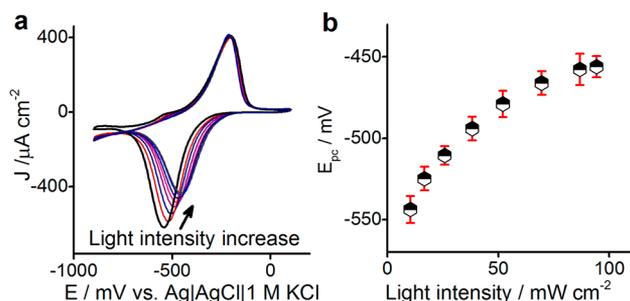


Figure 3. (a) Cyclic voltammetry of anthraquinone-terminated p-type silicon electrodes under different light intensities (10.3, 16.7, 25.8, 38.2, 52.0, 69.6, 86.7, and 94.2 mW cm^{-2}) in B&R buffer (pH 10.1). (b) Evolution of E_{pc} with light intensities. Illumination is operated on the whole surface from top side at the indicated intensity.

EIS experiments were performed under variable light intensity to investigate changes in the space charge layer as a result of the light and account for the apparent changes in the electron transfer kinetics. The Nyquist plots in Figure 4 exhibit

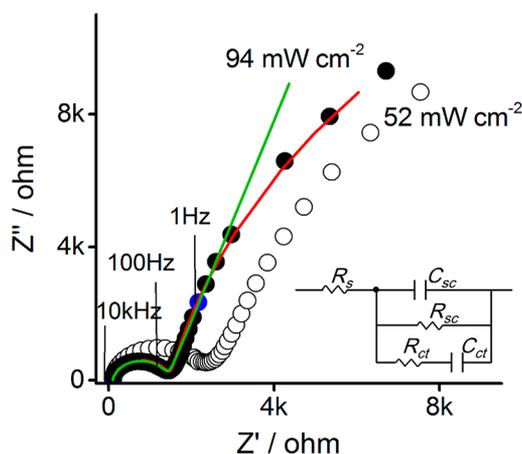


Figure 4. Nyquist plot for anthraquinone-terminated poorly doped p-type silicon at pH 10.06 under indicated light intensity. Circles are data from 100 kHz to 0.1 Hz, and red solid line is the best fit to the equivalent circuit shown in the inset, with a χ^2 statistics (R^2) of 0.001; the green solid line is the best fit to the simplified equivalent circuit by removing R_{sc} , with R^2 of 0.01.

a complete semicircle at the beginning and a second semicircle related to a slower process. The equivalent circuit used to refine the impedance data for the illuminated electrodes^{56,57} is shown in the inset. The semicircle in the high frequency region relates to the charge transfer process at the surface (R_{ct} and C_{dl}). The second semicircle arises from the space charge layer which is described by R_{sc} and C_{sc} . The series resistance is R_s . The Helmholtz layer was neglected in this case (infinite capacitance). The experimental R_{sc} and C_{sc} change with light intensity was shown in a logarithmic scale in Figure S6. With the increase of the illumination intensity there was an increase in C_{sc} and a slight decrease of R_{sc} .

3.5. Flat Band Potential Measurement by Mott–Schottky Plots. The flat band potential (E_{fb}) is an experimental parameter that can help defining the silicon band energy positions with respect to the potential of electroactive species in an electrolyte.⁵⁸ In the context of light activated electrochemistry, the knowledge of the value of the E_{fb} is vital as it indicates over which bias range the electrode

is expected to be in depletion—a key prerequisite for observing photoelectrochemical effects such as the light “switch on” of a faradaic process in a bias window where the depleted dark interface behaves like an ideally polarizable system (i.e., infinite resistance). As previously described,^{59–61} Mott–Schottky plots can be employed to determine E_{fb} from the relationship between the applied potential V and the space charge capacitance C_{sc} using the Mott–Schottky equation^{59–61}

$$C_{sc}^{-2} = \frac{2}{\epsilon \epsilon_0 q N A^2} \left(E - E_{fb} - \frac{kT}{q} \right) \quad (1)$$

where ϵ is the dielectric constant for silicon, ϵ_0 the vacuum permittivity, q the charge of the electron, A the electrode surface area, N the doping density of silicon, E the applied potential, and E_{fb} the flat band potential, and at room temperature kT/q is around 25 mV.

Figure S4 shows Mott–Schottky plots for the anthraquinone-terminated poorly doped p-type silicon at different pH values, from 5.06 to 10.09. In Mott–Schottky plots, linear regions were observed between -300 and -600 mV. The intersection point of the potential and linear $C^{-2}A^4$ potential curve gives a flat band potential, which in this case moved negatively with the increase of pH values, with a slope -22.5 mV pH^{-1} . The E_{fb} dependent on pH, which is also band edge shift with pH, reflects a minor amount of oxidation of silicon, below the resolution of the XPS, when contacted with aqueous. The pH sensitivity was consistent with previous pH response of $\text{CH}_2\text{CHCH}_2\text{-Si}$ -terminated silicon (-28 mV pH^{-1}) by Lewis and others.⁶² The carrier density ($(6.2 \pm 0.6) \times 10^{14} \text{ cm}^{-3}$) calculated from Mott–Schottky plots is consistent with the nominal resistivity values ($(6\text{--}12) \times 10^{14} \text{ cm}^{-3}$).⁶³

4. DISCUSSION

Here the p-type silicon was functionalized using the hydrosilylation reaction to form alkyl monolayer that protects the silicon against oxidation, followed by attaching an anthraquinone derivatives using a “click” reaction.^{20,36} Detailed characterization of XPS demonstrated that the anthraquinone moieties were successfully attached on silicon. As oxidation of the silicon surface results in an insulating layer, a modifying layer that prevents oxidation of the silicon in aqueous solution is of the major importance for developing silicon-based electrodes.^{36,42} The absence of silicon oxidation peaks in XPS spectra after the formation of the redox active self-assembled monolayer shows that this chemistry gives stable surfaces that will allow electrochemistry to be performed in the aqueous media without changes in the electrochemical performance of the electrode.⁴³

Highly doped silicon can be considered as having an impurity band which overlaps with the conduction or valence band. The overlap of the two bands results in free carriers such that the highly doped silicon acts more like a metal than as a semiconductor. Therefore, voltammograms of anthraquinone monolayers on highly doped silicon are very similar of those obtained on metallic electrodes. However, no redox behavior was observed at poorly doped silicon in the dark. Electrochemistry could be activated by illumination though. For poorly doped p-type silicon, the number of electrons in the conduction band is insufficient for the reduction of anthraquinone in the dark, and hence the faradaic processes were not observed in the cyclic voltammograms.^{3,7,12} However, electron–hole pairs generated inside the silicon after adsorption of light, and thus

an appreciable increase of electrons in the conduction band, activate the reduction to proceed, and hence the whole redox process of anthraquinone is observed. $E_{1/2}$ on poorly doped p-type silicon was altered to a more positive potential compared with the $E_{1/2}$ of anthraquinone on metal electrodes,^{29,51,52} which reflects the ability to reduce anthraquinone contra-thermodynamically by illumination. These results are consistent with the previous studies for the photoelectrochemical behavior of p-type silicon.^{2,3,7}

Generally, with a pH increase, the $E_{1/2}$ of anthraquinone moieties on metallic electrode will decrease according to the Nernst equation in the referred pH range (with a value around -59.2 mV per pH under standard conditions).²⁹ In the present work, the overall change of $E_{1/2}$ of anthraquinone on highly doped p-type silicon showed similar pH dependency with metallic electrodes. However, the anthraquinone-terminated poorly doped p-type silicon showed unexpected pH dependence on the redox behavior compared with the metal electrodes. The $E_{1/2}$ versus pH of anthraquinone for poorly doped p-type silicon was -44.0 mV pH^{-1} . Additionally, the pH dependences on E_{pa} and E_{pc} of anthraquinone on silicon electrode were quite different. The pH altered E_{pa} to a greater extent, while the linear dependency of E_{pc} on the pH value was very small.

The oxidation of hydroanthraquinone to anthraquinone involves loss of both electrons and protons. Together with the results from the Mott–Schottky plots, when the bias potential is more positive than flat band potentials, the silicon tends to be in accumulation. The oxidation of hydroanthraquinone could be considered as the injection of holes (valence band process) from the silicon. For p-type silicon electrode, the holes are in considerable excess of the electrons and are sufficient for the oxidation reaction to proceed. Hence, the deprotonation is probably the rate-limiting step. Supporting this notion is the positive shift in oxidation potential with a decrease in pH, which indicates the greater the amount of protons in the solution environment around the anthraquinone, the more energy that is required for the oxidation to proceed. Hence, pH alters E_{pa} to a greater extent. According to our previous work of the volumetric behavior on anthraquinone-terminated Au(111), there was a linear relationship between E_{pa} and pH and the slope of dependency of E_{pa} on the pH was -83.3 mV pH^{-1} .²⁹ The very similar dependency of E_{pa} with pH both for anthraquinone on gold electrode²⁹ and silicon electrodes supports the hypothesis that the deprotonation was the rate determining step for the oxidation of anthraquinone on silicon.

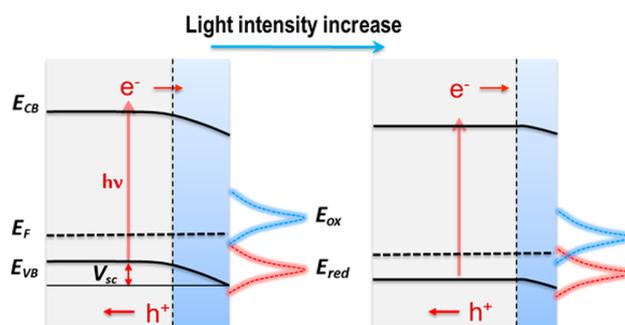
In contrast, the onset of the reduction process, E_{pc} , was insensitive to pH, hence suggesting the rate-determining step to be related to electron transfer rather than proton transfer. The reduction process of anthraquinone happens when the silicon is under depletion. The minority carrier (electrons) transfer has significant influence and pH appears to have little impact on the position of the reduction peak. Overall, the position of E_{pa} and E_{pc} can be controlled by deprotonation and electron transfer separately; as a result, the observed $E_{1/2}$ of the cyclic voltammograms did not follow the Nernst equation as it does on metal electrodes.

The difference between the conduction band potential and the anthraquinone reduction potential can be considered as the driving force for electron flow from silicon to the interface for anthraquinone reduction.^{48,62} The pH sensitivity of conduction band potentials correlates with flat band potential, which shifts negatively with a pH increase. As electron transfer is the

decisive factor for anthraquinone reduction on silicon, and the difference between the conduction band potential and the anthraquinone reduction potential was increased with pH, electron transfer would be thus facilitated. Meanwhile, for the oxidation process, deprotonation of the anthrahydroquinone would become easier with increases in pH values. The measured overall k_{et}' was found to increase with pH, which was consistent with the prediction as well as the kinetics change with pH on highly doped p-type silicon or metal electrodes.

The consequence of the hypothesis above that electron transfer is the rate-limiting step for the reduction process while proton transfer dominated for the oxidation process is that E_{pc} would be sensitive to light intensity while the E_{pa} will be insensitive to light intensity. This is in line with the experimental observations. The possible situations for the interface energetics in the present case are depicted in Scheme 2. The bands tend to be flatten with the light intensity

Scheme 2. Interface Energetics for Anthraquinone Derivatized Poorly Doped p-Type Silicon in Aqueous Electrolyte with Light Intensity Increase^a



^aThe symbols h^+ and e^- refer to holes and electrons, E_{CB} and E_{VB} stand for the conduction and valence band potentials of silicon, E_{pc} is the Fermi level of silicon, E_{ox} and E_{red} are the reduction and oxidation peak potentials of anthraquinone, and V_{sc} is the potential drop across the space charge layer. All other terms have their usual meaning.

increase,⁶⁴ and hence the Fermi level of silicon shifts, which leads to a change in electrode potential. This change was observed in Figure 3 where the reduction of anthraquinone became thermodynamically easier.

Our studies of the rate of electron transfer showed an increase in k_{et}' with increasing light intensity (Figure S5b). The dynamic processes at the silicon photoelectrode consists of minority carrier generation within the bulk, electron diffusion in the field free region, electron flow across the space charge layer, and charge transfer at the interface.⁶⁵ Since with the increase of illumination intensity more electrons will be generated in the conduction band of silicon and the electron diffusion range increases as well,^{66,67} impedance spectroscopy was used to explore changes in the space charge layer, and interface charge transfer resistance, as a function of illumination intensity. From the results shown in Figure S6a, C_{sc} was calculated to increase with light intensity. According to previous work, the relationship between the potential distribution in depletion layer, and its width, can be expressed by the Poisson–Boltzmann equation.^{5,6,65,67–69} The potential drop across the depletion layer (V_{sc} in Scheme 2) and its thickness would be smaller at higher light intensities. The C_{sc} shift is strong evidence for a decreased space charge layer width. At the same time, R_{sc} showed only a slight decrease with increased illumination

intensity. With a marked increase of electron flux in the bulk semiconductor, and the decrease of space charge layer thickness, the charge transfer would thus be facilitated at anthraquinone-terminated silicon interface. This is in agreement with a concomitant decrease of R_{ct} and increase of C_{ct} (Figure S6b). As a consequence, the apparent kinetics for electron transfer with higher light intensity were faster (Figure S5). Note however the decrease of R_{ct} with increased light intensity becomes less significant at higher light intensities, suggesting the electrons at the interface are becoming sufficient. In concert, the measured k_{et}' increased with light less significantly at higher light intensity.

5. CONCLUSIONS

In summary, an anthraquinone derivatized poorly doped p-type Si(100) system has been described, in which the $E_{1/2}$ of the redox species can be tuned either by changing the electrolyte pH or by varying the light intensity. The present work presents the significant pH dependency of the anodic process, leaving unaltered the cathodic event. Conversely, the position of the cathodic wave could be selectively tuned by changes in illumination intensity while the onset of the anodic event remains unaltered. The observation illustrates the oxidation and reduction state of anthraquinone relative to the energy bands of silicon. The oxidation process is a valence band process and holes transfer has negligible effects, while the reduction reaction relies on the electron transfer from the conduction band and the surface energetics has prominent influence. Both the anodic shift of reduction with respect to the silicon band level and increasing the intensity of illumination lead to easier electron transfer and hence facilitate the kinetics. The manipulation of silicon surfaces with anthraquinone and the possibility of tuning only the oxidation or reduction would serve to mediate the redox chemistry of solution species for further applications.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.jpcc.5b12097.

Figures S1–S6 and Table S1 (PDF)

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Notes

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