Synthesis and Characterisation of Chemical Bath Deposited TiO_2 Thin-Films

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ABSTRACT

A series of titania thin films was prepared by chemical bath deposition (CBD) of $TiCl_3$ on indium tin oxside (ITO) glass at room temperature, followed by calcinations at 500°C for 4 hours. The effect of cyclic deposition on phase composition, microstructure and electrical resistivity of TiO_2 thin films was characterised using X-ray diffraction, scanning electron microscopy and four-point probe respectively. Results showed that TiO_2 films produced by single deposition cycle were amorphous. In contrast, those produced by 5 and 6 deposition cycles were partly amorphous and partly crystalline with the formation of rutile. Both the film thickness and electrical resistivity increased with an increase in the number of deposition cycles.

Keywords: Chemical bath deposition, microstructure, rutile, amorphous, electrical resistivity.

1. Introduction

Thin-films are important components of various devices widely used in electronic, protective, membrane and sensor applications. The most common preparation method of thin-films is chemical deposition since it offers several advantages such as low cost, and the ability to produce samples with large surface area [1]. Some authors [1,2] have reported various aspects of chemically deposited thin films, while others [3,4] have studied the effects of varying the growth parameters such as deposition rates, bath compositions and bath temperature on the microstructures and properties of thin films.

TiO_2 exists in three main phases: anatase, brookite and rutile where the latter is a stable phase. This oxide is an attractive material due to its wide applications, such as in paints, sunscreens, and as photyocatalyst under ultraviolet light [5-7]. Hitherto, the synthesis of TiO_2 thin films or bulk has led to the emergence of nanoscience and nanotechnology by virtue of their superior physical and chemical properties [8,9].

In this paper, we have used chemical bath deposition to synthesize a series of titania thin films on indium tin oxide (ITO) glass and investigated the effect of cyclic deposition on phase composition, microstructure and electrical resistivity of TiO_2 thin films.

2. Experimental details

TiO_2 thin-films were synthesized using the solution of titanium trichloride ($TiCl_3$) and $NaHCO_3$. The process of deposition was carried out by dropwise addition of freshly prepared $NaHCO_3$ solution into $TiCl_3$ solution in a beaker glass, and then stirred until the pH of 3-4 was reached. The ITO glass substrate (35 x 15 x 1 mm^3) was hung vertically until 2 cm of the substrate was immersed into the solution of $TiCl_3$ and $NaHCO_3$. The solution was stirred for up to 6 hours, depending on the number of deposition cycles to produce 6 samples, as shown

in Table 1. The samples were then dried at 100^{o}C for 1 hr, followed by calcination at 500°C for 4 hr.

X-ray diffraction data were collected using a Shimadzu 610 diffractometer. The operating conditions used were CuKα radiation (λ = 1,54506 Å) produced at 40 kV and 30 mA. The diffraction patterns were collected over the 2θ range 10-80°. A JEOL JSM-6360LA scanning electron microscope was used to examine the microstructures of surface and cross-sections of the samples. The electrical resistivity (ρ) was of TiO_2 thin-films were measured using the four-point probe method.

3. Results and discussion

3.1 Phase composition

It was found that the color of TiO_2 thin-films is clear or transparent and adheres well to the ITO glass substrate. Visually the films are homogenous with smooth surface appearance. The XRD spectra were collected for samples A, D and F. The XRD data was also collected on ITO glass as a control. The diffraction pattern of sample F is shown in Fig.1 which depicts the existence of rutile only in the thin-film of TiO_2. In all samples, only rutile was detected and no peaks due to anatase (101) were detected probably due to their transformation to rutile during sintering at 500 °C for 4 h. Table 2 shows the phase analysis of anatase and rutile in all the samples prepared according to conditions stated in Table 1.

It should be mentioned that the normal XRD is not ideal for characterising the phase compositions of thin-films due to relatively high penetrating (≥30 μm) depth of x-rays in symmetric mode. Since the TiO_2 thin-films produced in this study are less than 5 μm thick, the x-rays will penetrate into the ITO substrate and produce an undesirable strong background noise due to the amorphous nature of ITO. Hence, it will be more accurate to use grazing-

incidence XRD to probe their near-surface phase compositions of TiO_2 thin-films. The critical angle (α_c) for grazing-incidence diffraction can be approximated by [10]:

$$\alpha_c = 1.6 \times 10^{-3} \rho \lambda$$

where α_c is in radians, ρ is density of TiO_2 (rutile = 4.25 g/cm^3) and λ is wavelength in Å. Hence the calculated value of α_c for TiO_2 thin films is 0.010472 or 0.6°. Above the critical angle, the penetration depth (d) can be calculted from [11]:

$$d = 2 \alpha/\mu$$

where μ is the linier absorption coefficient of matter. Since the value of μ for rutile is 907.0775/cm [12] and the normal XRD uses an incidence angle of 8°, penetration depth of x-rays will be ~30.8 μm which is deeper than the thickness of the thin-films. As a result, the x-rays will also detect the ITO glass and give rise to undesirable large amorphous background.

By using the Scherrer equation [13], one can determine the crystallite size of this sample. By taking the width of diffraction peak into consideration, the crystallite size of TiO_2 in sample F was determined to be ~10 nm. This value is much less than the particle size of ~ 200 nm indicated by the SEM image (Fig. 2d). Since particles or grains are made up of small crystallites, this suggests that there are about 20 crystallites in each TiO_2 particle on sample F.

The PDF file number 21-1276 for rutile from the database was used to identify the phases formed [14]. According to Lokhande [15], there is no difference between ITO glass and TiO_2 having low grade of crystallinity. In fact, it is not necessary for TiO_2 to have a high grade of crystallinity when it is used in applications such as a photodetector [16].

3.2 Microstructure

The SEM images of TiO_2 thin films prepared by multiple depositions are shown in Fig. 2. These images clearly show distinct contrast in microstructures in films prepared between one

and six deposition cycles. This indicates that the number of deposition cycles can lead to a significant change in the microstructures of TiO_2 films.

The microstructures of the thin-films shown in Fig. 2 are for samples with 1, 2, 4 and 6 deposition cycles. It is interesting to note that the microstructures for thin-films prepared from 1 to 4 cycles are quite similar although cracking can be seen in sample A with one deposition cycle. In films prepared from 4 and 6 deposition cycles, the cracks had disappeared and the amount of pores also decreased. It appears that multiple depositions had helped to fill the pores within the TiO_2 layer thus making the surface denser. If the film remains attached to the substrate and does not crack during sintering, shrinkage in the plane of the substrate can be inhibited. The stresses that arise during the sintering of constrained films are analogous to those in a sandwich seal caused by mismatch in the thermal expansion coefficients in the layers.

It is interesting to note that the films have pores distributed all over the surfaces. These pores can play the role of enhancing dye-sensitization when the TiO_2 films are used for solar cell application [17] as the dye can fill the pores. Dye can also absorb the incident ray and inject electrons to the TiO_2 semiconductor.

Furthemore, samples were also observed using SEM in cross-sectional view from top surface of the film to the ITO glass in order to estimate the film thickness. Microstructures of samples A, B and E are shown in Fig. 3. Using the scale-bar on each image, one can estimete the film thickness deposited on the substrate. The variation of film thickness as a function of deposition cycles is depicted in Fig. 4. By taking into account the two estimated standard deviations, the thickness of the films can be observed to increase with the number of deposition cycles. This is self-evident because when repeated depositions were done. the material deposited on the substrate also increased. From this figure one can tailor-design the thickness of the film based on number deposition cycles.

3.3 Electrical Resistivity

The electrical resistivity or conductivity of TiO_2 films was measured using four-point probes. By knowing the resistivity one can infer whether film produced is a conductor or semiconductor. The application of TiO_2 as photo-conducter must have properties of a semiconductor because a conductor has no band-gap energy. If the incident ray overcomes the band-gap energy, electrical conduction will occur. The variation of electrical resistivity as a function of deposition cycles is depicted in Fig. 5.

The resistivities of conductors and semiconductors normally range from 10^{-3} -10^{8} Ω-cm [18]. Based on this value, the resistivity of TiO_2 thin-films synthesized in this study can be classified as a semiconductor. Since the resistivity increases with deposition cycles, this indicates that multiple depositions serve to increase the resistivity by virtue of more pores and materials being deposited in the film as the thickness increases. When the deposited material increases, the resistivity of a layer increases with each deposition cycle. It has been reported that the resistivity value of TiO_2 thin films prepared by DC magnetron sputtering ranged from 10^{7}-10^{10} Ω-cm [19]. Thus, the resistivity results obtained in this work are comparable to results reported in the literature.

4. Conclusions

Thin-films of TiO_2 produced by multiple depositions were more crystalline with rutile being the only phase formed. In contrast, TiO_2 films produced by single deposition cycle were amorphous. Both the film thickness and electrical resistivity of TiO_2 increased with an increase in the number of deposition cycles.

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Figure Captions

1. The XRD plot for sample F with six cycles of deposition. [Legend: r = rutile]
2. SEM images of samples with various processing conditions: (a) sample A (1 cycle @ 6 hr), (b) sample B (2 cycles @ 3 hr), (c) sample D (4 cycles @ 1.5 hr), and (d) sample F (6 cycles @ 1 hr).
3. SEM images showing the cross-sectional view from of samples from top-surface to the ITO glass substrate where the TiO_2 layer is on the left and the ITO glass is on the right. (a) sample A, (b) sample B and (c) sample E.
4. Variations of film thickness as a function of deposition cycle.
5. Variations of electrical resistivity of TiO_2 films as a function of deposition cycle.