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Large-scale and Rapid Synthesis of
Disk-Shaped and Nano-Sized GrapheneChunyong He¹, San Ping Jiang² & Pei Kang Shen¹

SUBJECT AREAS:

MATERIALS SCIENCE

NANOSCIENCE AND
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PHOTOCATALYSIS

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¹State Key Laboratory of Optoelectronic Materials and Technologies, and Guangdong Province Key Laboratory of Low-carbon Chemistry & Energy Conservation, School of Physics and Engineering, Sun Yat-sen University, Guangzhou, 510275, PR China, ²Fuels and Energy Technology Institute & Department of Chemical Engineering, Curtin University, Perth, WA6102, Australia.

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We synthesized disk-shaped and nano-sized graphene (DSNG) through a novel ion-exchange methodology. This new methodology is achieved by constructing metal ion/ion-exchange resin framework. The morphology and size of the graphene can be modulated by changing the mass ratio of the carbon-containing resin to the cobalt-containing precursor. This is the first time to show that the DSNG formed on the granular transition metal substrate. The DSNG gives a high intensity of photoluminescence at near-UV wavelength of 311 nm which may provide a new type of fluorescence for applications in laser devices, ultraviolet detector UV-shielding agent and energy technology. The emission intensity of the DSNG is thirty times higher than that of the commercial large graphene. Our approach for graphene growth is conveniently controllable, easy to scale-up and the DSNG shows superior luminescent properties as compared to conventional large graphene.

Correspondence and requests for materials should be addressed to P.K.S. (stsspk@mail.sysu.edu.cn) or S.P.J. (s.jiang@curtin.edu.au)

It is well-known from many theoretical and experimental evidences that nanographene with specifically morphologies have excellent properties for broad application prospects^{1–4}. The applications of graphene have been widely studied^{4–6}. The amazing properties of graphene make it a likely candidate to spur on the next revolution in electronics⁷. However, large-area graphene flakes are not semiconductor because there is no energy gap. But its energy gap can be opened when graphene is shaped up into quasi-one-dimensional structures with narrow widths (<~30 nm), and similar to CNTs, the energy gap of graphene is tunable from 0 to 0.5 eV, depending on the width and crystallographic orientation of the graphene nanoribbon (GNR)^{2,8}. Graphene also touted as the basis for next generation spintronics devices^{9,10}. Arbitrarily shaped finite graphene nanoflakes (GNFs) can lead to large net spin or antiferromagnetic coupling between groups of electron spins. And the spin of GNFs depends on its shape due to topological frustration of the δ -bonds^{11–13}. Large area (>100 nm) graphene behaves as conventional single-electron transistors, exhibiting periodic Coulomb blockade peaks. When the size of graphene narrows down to 100 nm, even sub-10 nm, which refer as graphene quantum dots (GQDs), the peaks become strongly nonperiodic, indicating a major contribution of quantum confinement¹⁴. The GQDs with sub-10 nm in width remain conductive and reveal a confinement gap of up to 0.8 eV, demonstrating the possibility of molecular-scale electronics based on graphene^{15,16}. Apart from potential applications for the top-down approach to molecular-scale electronics, GQDs also have other unique properties and potential applications in many fields, such as photovoltaic devices⁴, biological label¹⁷, light emitting diodes¹⁸, and sensor¹⁹.

Although various effective techniques, such as mechanical exfoliation¹, oxidation of graphite²⁰, liquid-phase exfoliation²¹, epitaxial growth²² and thermal decomposition of SiC²³ have been developed to produce single-layer or few-layers graphene, the controlled method of preparing nanographene still counts for the top issue²⁴. Han *et al.*² used lithography etching technology lithographically patterned graphene nanoribbon (GNR) structures. Tapaszo *et al.*⁸ successfully obtained well-defined nanoribbons with widths down to 2.5 nm by applying a constant bias potential (significantly higher than the one used for imaging) and simultaneously moving the STM tip with constant velocity. Campos *et al.*²⁵ demonstrated anisotropic etching of the single-layer graphene by thermally activated nickel nanoparticles and obtain sub-10-nm nanoribbons. Dai *et al.*²⁶ developed a route in which multi-walled carbon nanotubes were partially embedded in a PMMA film and then etched the exposed tube walls by argon plasma. The resulting nanoribbons have smooth edges and very narrow width (10–20 nm). Pan *et al.*²⁷ have developed a hydrothermal route for cutting preoxidized graphene sheets into GQDs (9.6 nm in average diameter) with blue emissions. Lu *et al.*¹⁵ demonstrated the formation of regularly sized graphene quantum dots on Ru (0001) substrate by using C₆₀, which acts as a unique precursor compared to C₂H₄. Zhuo *et al.*²⁸ presented a new facile ultrasonic method to prepare GQDs, which exhibited bright PL in a water solution.

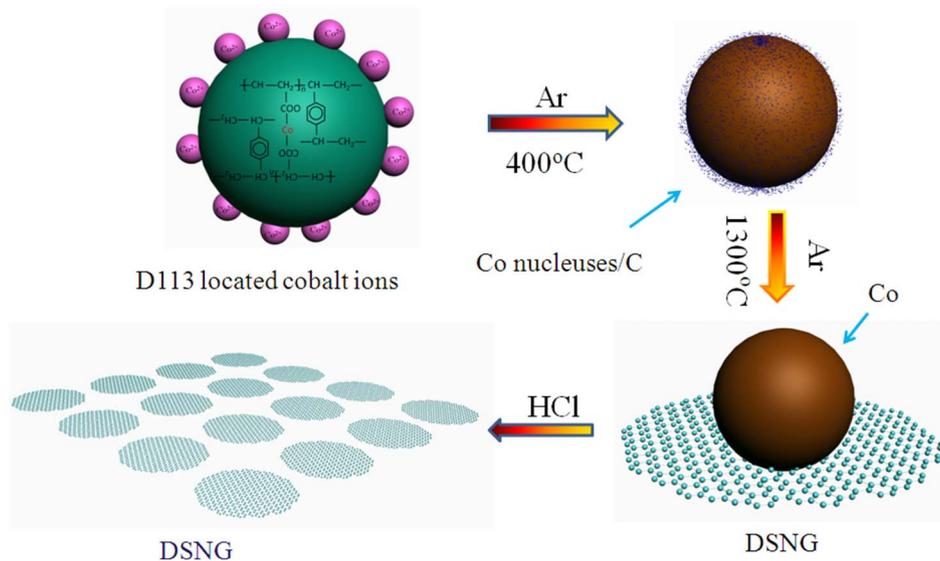


Figure 1 | Schematic view of the approach for disk-shaped and nano-sized graphene (DSNG) growth.

To our best knowledge, the disk-shaped and nano-sized graphene (DSNG) has rarely been reported. The arduousness of obtaining DSNG with well-defined shape is the main reason of the poverty in study. Although some techniques, such as chemically derived techniques²⁹, chemical vapor deposition on template³⁰ and lithography etching³¹ have been reported, however, effective method to produce DSNG with well-defined shape remains elusive. The disk-shaped and nano-sized structure makes the DSNG possess much higher edge density and fringe defects than that of normal graphene sheets, which will contribute to better edge reactivity and defects-mediated properties^{31,32}. The structural defects may give rise to localized electronic states and lead to the magnetic state in DSNG^{32–34}. These can offer key advantages in realizing various electron applications via edge chemical functionalization, such as doping.

Recently, graphene based nanomaterial³⁵ are attracting more and more attentions as luminescent nanomaterials due to their unique properties, such as low toxicity, high luminescent, robust chemical inertness and environmentally-friendly.

In this paper, we directly synthesized DSNG through a whole new methodology. The new methodology is achieved by constructing metal ion-ion exchange resin framework. The high quality of DSNG was confirmed by atomic force microscopy (AFM), scanning electron microscopy (SEM), Raman, X-ray Diffraction (XRD) and X-ray photoelectron spectroscopy (XPS). Our approach for DSNG growth is actually utmost simple, conveniently controllable and easily scale-up. The synthetic process of the DSNG is shown in Figure 1.

Results

The AFM was used to characterize the material deposited on new-delaminated mica substrate at ambient conditions. It is observed that the material is disk-shaped with the diameter around 200 nm (Figure 2a). The DSNG gave the appearance of the uniform disks. To obtain the morphology and structure of the DSNG, we measured the thickness of the sample by AFM. Figure 2b shows the AFM result that the nano-sized graphene were disk-shaped with the height of 0.74 nm. Figure 2c displays the scanning electron microscope (SEM) of the as-grown DSNG which are well dispersed. As shown in Figure 2d, a perfect disk-like shape can be clearly seen in the SEM micrograph with the size around 200 nm.

The D113 resin exchanged with cobalt ions was heated in tube furnace at 400°C under Ar atmosphere for 2 h, resulting in

amorphous powder which was confirmed by the XRD measurements (Figure S3, pattern a). This step is essential in the synthetic process of DSNG, otherwise, no inerratic DSNG is observed (Figure S2). When the sample was further heat treated at 1300°C, the polycrystalline cobalt particles and graphitic carbon formed (Figure S3, pattern b). The Co was removed by HCl treatment for 12 h (Figure S3, pattern c).

The Raman spectra of the HOPG and DSNG in the range of 1000–3000 cm^{-1} are shown in Figure 3a. The band observed at about 1577 cm^{-1} is G band, which corresponds to a splitting of the E_{2g} stretching mode of graphite and reflects the structural intensity of the sp^2 -hybridized carbon atom. The 1350 cm^{-1} band (D band), assigned to zone center phonons of K-point phonons of A_{1g} symmetry, characterizes the disordered graphite planes and the defects incorporated into pentagon and heptagon graphitic structures. The minor D band (Figure 3a) in DSNG mainly due to the huge amounts

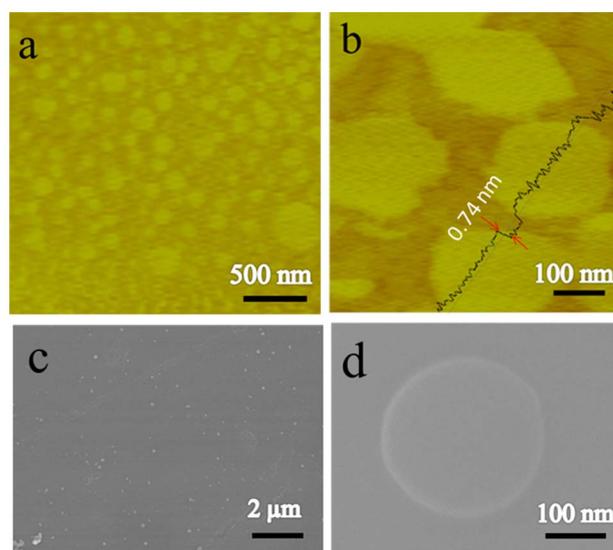


Figure 2 | Microscopic images of DSNG. (a). The tapping mode AFM image of the disk-shaped nano-sized graphene (DSNG) prepared at mass ratio of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ to D113 resin ($R_{\text{Co/R}}$) of 0.1, (b). a close-up of DSNG, (c). SEM micrograph of DSNG and (d). high resolution SEM micrograph of a DSNG.

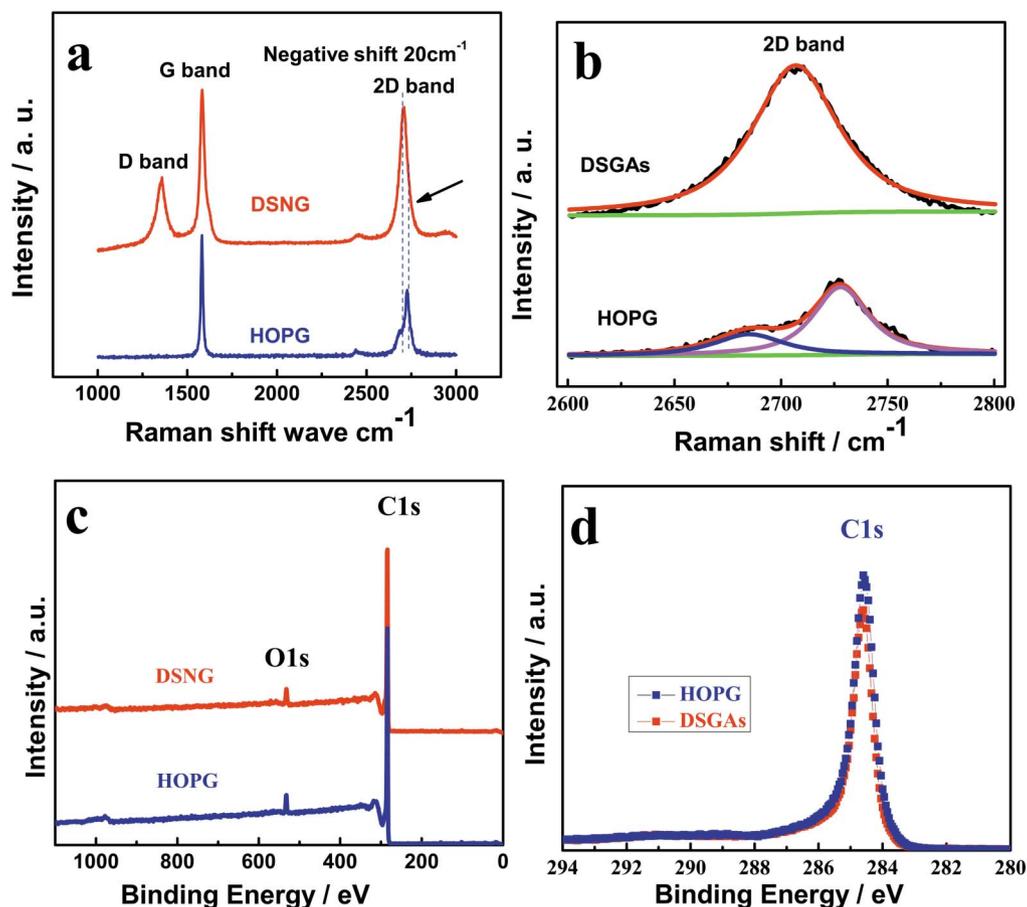


Figure 3 | Raman and XPS spectra of DSNG. (a). Comparison of Raman spectra of DSNG (red line) and highly oriented pyrolytic graphite (HOPG) (blue line) at 514 nm, (b). the 2D band in the Raman spectrum obtained from DSNG is formed by one Lorentzian peak, and the HOPG is formed by a two-peak line shape, (c). comparison of XPS spectra of DSNG (red line) and HOPG (blue line) and (d). high-resolution C1s XPS spectra of DSNG (red line) and HOPG (blue line).

of edges³⁶. The result shows that a significant change in shape and intensity of the 2D band of DSNG compared to HOPG (Figure 3b). The 2D band in the Raman spectrum obtained from HOPG is formed by a two-peak line shape, consisting of two components $2D_1$ and $2D_2$ ³⁷. The 2D band in the Raman spectrum of DSNG is a single and sharp peak, formed by one Lorentzian peak³⁸. Notably, the intensities of the 2D band and G band for DSNG are comparable. The 2D position of DSNG is 20 cm^{-1} shift to lower frequency than that of the HOPG.

X-ray photoelectron spectroscopy (XPS) is a powerful spectroscopic technique that measures the elemental composition, empirical formula, chemical state and electronic state of the elements existed in a material. XPS was used to study the bonding state of carbon atom in the DSNG. The full scan spectra of the DSNG are shown in Figure 3c. The peaks at binding energy of 284.6 eV and 533.15 eV are ascribed to C1s and O1s. The O1s is ascribed to the adsorbed molecular oxygen on the surface of the DSNG and HOPG. The hybridization of carbons influences the properties of a carbon nanomaterial and has implications in its structure. XPS can be used to determine the hybridization of the carbons on the surface of a material, such as graphite and nanodiamond. Graphite is a carbon material consisting of sp^2 carbons. Theoretically the XPS of the pure graphite will show a single C1s peak, with a binding energy characteristic of sp^2 carbon. Figure 3d shows the detailed C1s XPS spectrum of DSNG (red line) and HOPG (blue line). The main peak at 284.6 eV corresponds to the graphite-like sp^2 bonded carbon³⁹. The C1s spectrum of DSNG coincides with HOPG, indicating most of the carbon atoms in DSNG are arranged in a conjugated honeycomb lattice³⁹.

Figure 4 demonstrates that the morphology of the samples is a function of mass ratio of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ to D113 resin ($R_{\text{Co/R}}$). It was found that the samples display different geography and size when the $R_{\text{Co/R}}$ changes. In the case of the $R_{\text{Co/R}} = 0.2$ and 0.5, there are two kinds of DSNG coexisted with different dimensions of the samples (Figure 4a and 4b). One is bigger and another one is very small. When the $R_{\text{Co/R}} = 1$, the geography of the graphene is inclined to become random (Figure 4c). The AFM height analysis illustrates that the heights of all the samples are between 0.45 nm to 0.71 nm (see

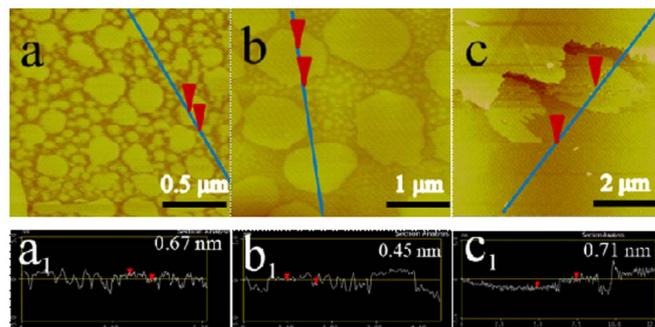


Figure 4 | AFM images of the graphene samples. The tapping mode AFM images of the graphene samples deposited on new-delaminated mica at ambient condition as a function of mass ratio of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ to D113 resin: (a). $R_{\text{Co/R}} = 0.2$, (b). $R_{\text{Co/R}} = 0.5$ and (c). $R_{\text{Co/R}} = 1$. The a_1 – c_1 are the height profiles along the blue lines in (a–c), respectively.

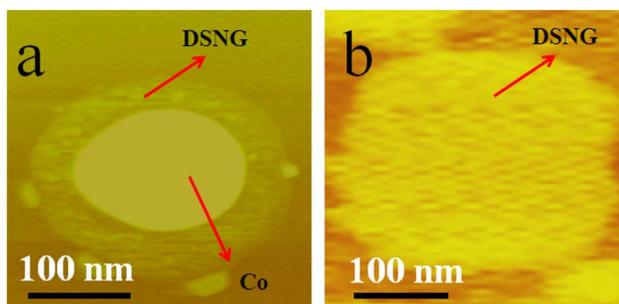


Figure 5 | AFM images of DSNG. (a). The tapping mode AFM image of the Co/graphene before acid treatment ($R_{Co/R} = 0.1$) and (b), the tapping mode AFM image of DSNG after acid treatment ($R_{Co/R} = 0.1$).

Figure 4a₁–c₁). The Raman spectra (Figure S4) demonstrate that all the samples possess the characteristics of graphene. The XPS spectra (Figure S5) of the samples show that the carbon atoms in the samples are sp^2 hybridized carbon in the regular hexagon plane, respectively.

The DSNG is formed by the catalysis of cobalt during the heat treatment which is confirmed by the AFM measurements before and after acid treatment. Figure 5 shows the typical results on a sample prepared in the case of the $R_{Co/R}$ value of 0.1. Figure 5a presents a disk there a Co particle located on the surface of graphene. The peripheral graphene disk is lighter yellow and the Co particle gives darker color, respectively. A flat disk formed after the removal of the Co particle as shown in Figure 5b. The AFM images of the Co on graphene samples with $R_{Co/R} = 0.1, 0.2, 0.5$ and 1 are shown in Figure S6 and S7.

Highly efficient luminescent materials hold the key for applications in optoelectronic devices, biological labeling and sensing^{40–42}. Carbon-based nano-materials have been attracting the leading attention as they show high efficiency, stable emissions, chemical inertia,

nontoxicity in the physiological condition and no environmental pollution⁴³. With this in mind, we recorded the photoluminescent excitation (PLE) spectrum and PL emission spectrum of DSNG, and the large-area graphene was also tested. A PL emission with high intensity observed from DSNG shows a near-UV wavelength at 311 nm (3.98 eV) under an excitation wavelength of 270 nm (4.62 eV) (Figure 6b). The emission intensity of DSNG is thirty times as high as that of large-area graphene. The enhanced emission indicates that there are more emission active sites. Thus, we expect that the nanostructured DSNG of sp^2 configuration of similar size is likely to be responsible for the enhanced near-UV PL. The PLE spectrum for DSNG is shown in Figure 6a. Excitonic features are readily observable between excitation wavelengths of 240 and 295 nm, which represent the absorption energies corresponding to emission of near-UV⁴⁴. Figure S8 shows the Photograph of suspension of DSNG (1 mg ml^{-1}) in H_2O , N-methyl-2-pyrrolidone (NMP), N,N-Dimethylformamide (DMF) and 1,2-Dichloroethane (EDC).

The PL performance of functionalized graphene is related with the size, shape, defects, and preparation method. Table S1 displays different examples of the photoluminescence from functionalized graphene. Figure 6c compares the PL for the graphene samples derived from different $R_{Co/R}$. All the peaks are in the same near-UV. The emission intensity increases with the decreasing of $R_{Co/R}$ (Figure 6d). The origin of photoluminescence (PL) in these carbon-based nano-materials is tentatively proposed to be from isolated polyaromatic structures or passivated surface defects⁴³. These results show that the state of the active emission point of the samples are the same. The increase in the emission intensity is mainly due to the increase in the fringe defects which result from the decrease of graphene sheet size. Figure S9 shows the relationship between the normalized PL intensity and the graphene size. The emission intensity decreases with the increase of the graphene size.

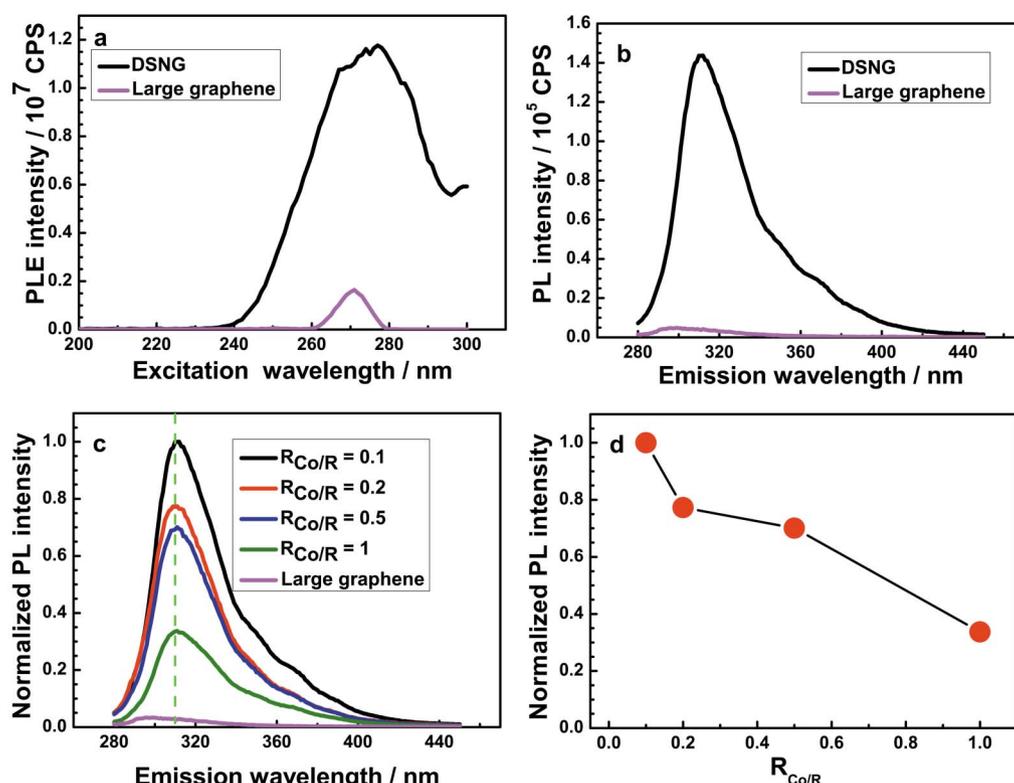


Figure 6 | PL spectra of DSNG. (a). PLE spectra of the DSNG ($R_{Co/R} = 0.1$) and large-area graphene with the detection wavelength of 311 nm and (b). PL spectra of the DSNG ($R_{Co/R} = 0.1$) and large-area graphene excited at 270 nm, (c). PL spectra of the nanographene samples with different $R_{Co/R}$ and (d), the relationship between the normalized PL intensity and the $R_{Co/R}$.



Discussion

The synthesis of large area, homogenous graphene on cobalt by chemical vapor deposition CVD method has been reported^{45–48}. The synthetic mechanism of cobalt catalyzed graphene by CVD is shown as follows. A solid solution of cobalt and carbon is formed when the mixture of H₂ and amorphous gaseous carbon (a-C) source is introduced to the hot cobalt film. In the process of cooling down, the high energy carbon atoms in cobalt metal overcome the energy barrier and segregate from the interior of the cobalt film to the surface to form graphene^{45,47}. The graphene grown by CVD on cobalt substrate is extensive, low-defect and homogeneous. Our approach for DSNG growth is schematically illustrated in Figure 1. The D113 resin exchanged with metal ions were carbonized in tubular furnace at 400 °C under Ar atmosphere for 2 h. In this process, the cobalt atoms, encapsulated in the exchange sites of the D113 resin, bring together into nucleuses. Simultaneously, the D113 resin began decomposition. After going through a carbonization process, the a-C, derived from the carbonization of the D113 resin, adsorb on the surface of the cobalt nucleuses. The compound is still amorphous solids (Figure S3 pattern a). Next, the compound was then heated to 1300 °C for 1 h under Ar atmosphere condition. At this stage, the cobalt nucleuses grow and become polycrystalline cobalt particles with the increase of the temperature (Figure S3, pattern b). The sample was then cooled down to room temperature. In the process of cooling, the carbon atoms diffuse out of the cobalt to the surface, forming the nanographene. These infant nanographenes are expanding outward homogeneously at all orientations, forming a series of inerratic DSNG (Figure 5a).

In conclusion, we reported a novel ion-exchange method to synthesize the DSNG with superior luminescent properties. This is the first example to show that the DSNG formed under the catalysis of the granular cobalt. The morphology and size of the graphene can be modulated by changing the mass ratio of the cobalt-containing precursor to the carbon-containing resin. The DSNG synthesized in this work showed high efficiency in near-UV luminescence which may provide a new type of fluorescence for applications in laser devices, ultraviolet detector UV-shielding agent and energy technology. They may expand the application of graphene-based materials to other fields. Our approach for graphene growth is conveniently controllable, easy to scale-up and readily for applications.

Methods

Preparation of disk-shaped and nano-sized graphene. In a typical preparation, 2.0 g CoCl₂·6H₂O was dissolved in 200 ml deionized water and 20 g D113 resin was put into the solution (the mass ratio of CoCl₂·6H₂O to D113 resin: R_{Co/R} = 0.1). After magnetic stirring for 2 h, the D113 resin loaded with metal ions were rinsed with deionized water for several times and dried in oven at 80 °C. Figure S1 shows the typical profile of the ion exchange process of D113 between Co₂⁺ and H⁺. After carbonized in a tube furnace at 400 °C under Ar atmosphere condition for 2 h, the samples were further heat treated in tube furnace at 1300 °C for 1 h. The product was post-treated in 3.0 mol L⁻¹ HCl solution for more than 12 h with magnetic stirring to remove cobalt on the surface of the graphene.

The other graphene samples with different morphologies are synthesized by the same steps as mentioned above. In these preparations, 4.0 g, 8.0 g, 20 g CoCl₂·6H₂O was dissolved in 200 ml deionized water and 20 g D113 resin was put into the solution (R_{Co/R} = 0.2, 0.5 and 1), respectively. After treated by the same steps mentioned above. The graphene samples with different morphologies were obtained.

More details on the experimental section are presented in the Supplementary Information.

Characterizations. The X-ray diffraction (XRD) analysis was performed on a D/Max-III (Rigaku Co., Japan) using Cu K α radiation, and operating at 40 kV and 30 mA. The 2 θ angular regions between 10° and 110° were explored at a scan rate of 5° min⁻¹. The Raman spectroscopic measurements were carried out on a Raman spectrometer (Renishaw Corp., UK) using a He/Ne laser with a wave length of 514.5 nm. The X-ray photoelectron spectroscopy (XPS) measurements were carried out on a XPS apparatus (ESCALAB 250, Thermo-VG Scientific Ltd.). A field emission scanning electron microscope (FESEM, JSM-6330F, JEOL Ltd.) and multimode scanning probe microscope (SPM, Nanoscop IIIa, Digital Instrument Ltd.). PL spectra were measured at room temperature in a FLS920P system (EDINBURGH INSTRUMENTS Ltd.) with a 450 mW xenon arc lamp. All the samples were diluted

with DI water to avoid reabsorption; the samples were placed in the cuvettes inside the integrating sphere.

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Author contributions

C.Y.H. and P.K.S. conceived the experiments. C.Y.H. performed the experiments. C.Y.H., P.K.S. and S.P.J. wrote the paper. All authors contributed through scientific discussions.

Additional information

Supplementary information accompanies this paper at <http://www.nature.com/scientificreports>

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